

Predicting Products of Chemical Reactions

This worksheet is designed to help you predict products of simple reactions of the four basic reaction types (synthesis, decomposition, single replacement, and double replacement) and combustion reactions.

For the first few reactions, the type of reaction is listed, you should predict the products, then balance. Further questions just have the reactants listed and you should decide on the type of reaction, as well as the correct products. Many of these reactions fall into the category of redox reactions, though do not let that confuse you...each can be described in terms of the four basic reaction types (except the combustion reactions).

Although states (s, l, g, aq) of the reactants and products are very important in a chemical reaction, don't worry about determining those for these problems. Rather, focus on what products might result from the reactants given. Pay particular attention to the ionic charge of species that you know forms ions with only one possible charge (e.g., alkali metals, alkaline earth metals, halogens, etc.)

1. Combustion: $\text{C}_2\text{H}_6 + 9 \text{O}_2 \rightarrow 6 \text{CO}_2 + 6 \text{H}_2\text{O}$
2. Combustion: $2 \text{C}_4\text{H}_{10} + 11 \text{O}_2 \rightarrow 8 \text{CO}_2 + 6 \text{H}_2\text{O}$
3. Combustion: $\text{C}_5\text{H}_{12}\text{O}_2 + 7 \text{O}_2 \rightarrow 6 \text{CO}_2 + 5 \text{H}_2\text{O}$
4. Synthesis: $\text{Mg} + \text{I}_2 \rightarrow \text{MgI}_2$
5. Double displacement: $\text{CuCl}_2 + \text{H}_2\text{S} \rightarrow \text{CuS} + 2 \text{HCl}$
6. Double displacement: $\text{NaOH} + \text{HClO}_4 \rightarrow \text{NaClO}_4 + \text{H}_2\text{O}$
7. Decomposition: $\text{ZnCO}_3 + \text{heat} \rightarrow \text{ZnO} + \text{CO}_2$
8. Single replacement: $2 \text{HCl} + \text{Zn} \rightarrow \text{ZnCl}_2 + \text{H}_2$
9. single $2 \text{Na} + \text{MgCl}_2 \rightarrow 2 \text{NaCl} + \text{Mg}$
10. double $\text{CaCl}_2 + \text{K}_2\text{CO}_3 \rightarrow \text{CaCO}_3 + 2 \text{KCl}$
11. synthesis $2 \text{K} + \text{Cl}_2 \rightarrow 2 \text{KCl}$
12. double $3 \text{BaCl}_2 + 2 \text{K}_3\text{PO}_4 \rightarrow \text{Ba}_3(\text{PO}_4)_2 + 6 \text{KCl}$
13. double $\text{H}_2\text{SO}_4 + 2 \text{KOH} \rightarrow \text{K}_2\text{SO}_4 + 2 \text{H}_2\text{O}$
14. synthesis $4 \text{Al} + 3 \text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$
15. double $\text{Pb}(\text{NO}_3)_2 + 2 \text{KOH} \rightarrow \text{Pb}(\text{OH})_2 + 2 \text{KNO}_3$
16. double $\text{H}_2\text{SO}_4 + \text{BaCl}_2 \rightarrow 2 \text{HCl} + \text{BaSO}_4$
17. single $\text{Ca} + 2 \text{AgCl} \rightarrow \text{CaCl}_2 + 2 \text{Ag}$
18. double $\text{H}_3\text{PO}_4 + \text{FeBr}_3 \rightarrow 3 \text{HBr} + \text{FePO}_4$
19. synthesis $6 \text{Li} + \text{N}_2 \rightarrow 2 \text{Li}_3\text{N}$
20. double $2 \text{HCl} + \text{Mg}(\text{OH})_2 \rightarrow \text{MgCl}_2 + 2 \text{H}_2\text{O}$

Predicting Products And Balancing Equations Answers

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Khomtchouk,2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields Chemical Innovations:

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