

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.
Calculate: Molar masses, stoichiometric ratios, limiting reagent, percent yield, theoretical yield, actual yield, percent error.

Chemists will often make reactions in which one reactant is in excess and the other is limiting. The limiting reagent is the one that runs out first.



The ratio required is 2:1 (molar) for hydrogen to oxygen.

2H ₂ (g)	20.0 g/mol
O ₂ (g)	32.0 g/mol
2H ₂ O(l)	36.0 g/mol
H ₂ O	18.0 g/mol

1. 20.0 g of each element will be used. How much H₂O can be produced from 20.0 g of each?
 Remember: Limiting reagent will be the one that runs out first.



2. Which is the limiting reagent? (The one that runs out first.)

3. How much H₂O can be produced from 20.0 g of each element? (The limiting reagent will be the one that runs out first.)



4. If the reaction is performed in a lab, 22.5 g of H₂O is produced. What is the percent yield?



Percent Yield And Limiting Reactants Quiz

Purshotam Kaushik & Kirti Kaushik



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