

A. Identify an equilibrium, the chemical equation, the balanced chemical equation, and the equilibrium constant expression.



Classroom: A student asks you where the K_{eq} is. You say it is in the reaction. The student then asks you if the K_{eq} is always the same for a given reaction. How do you respond?

A. It is always the same.

B. It is the same for a given reaction, but it is different for different reactions. (This is the correct answer.)

C. It is the same for a given reaction, but it is different for different reactions. (This is the correct answer.)



$$K_{eq} = \frac{[C][D]}{[A][B]}$$

For K_{eq} in the reaction, the product is the product.

Product	Reactant
Reactant	Product
Product	Reactant

Ex. 10: Calculate K_{eq} for the following reaction using the given equilibrium concentrations. Then determine whether the forward or reverse reaction is favored.



$$\frac{[H_2O(l)][CO_2(g)]}{[H_2O(g)][CO_2(g)]}$$

$$\frac{[0.001][0.001]}{[0.001][0.001]}$$

$$K_{eq} = 0.001$$

Note Taking Guide Episode 1202 Answers

Ying-Ying Zheng



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