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# ACS Gen Chem 1 Exam

**Ion** - A species of an element in which the number of electrons does not equal the number of protons.

**Isotope** - Each isotope of an atom has a different number of neutrons.

**Isoelectric** - Two atoms with the same charge.

**Isotopic** - Two atoms with the same number of neutrons.

**Transition Metal Orbitals** - The 4s orbital is filled before the 3d orbital, but listed after.

**Paramagnetic** - Atoms that possess a permanent magnetic charge due to the presence of unpaired electrons.

**Electropositive** - Tending to lose electrons and form positive ions.

**Tetrahedral Bond Angle** -  $109.5^\circ$

**Resonance Structures** - Two equally correct arrangements of one molecule's electrons; generally the molecule exists as an average of the two.

**Polar Molecule** - A molecule that is mostly positive on one side and mostly negative on the other.

**$sp^3$  Orbital** - An orbital composed of one s orbital and three p orbitals; can hold eight total electrons. Tetrahedral.

**Trigonal Planar** - The way that three objects will arrange themselves around a central atom. Non-polar.

**Free Radicals** - Compounds with an odd number of electrons in their Lewis Structure. Eg.  $\text{NO}_2$

**Incomplete Octets** - Occur in elements that can form stable bonds with less than eight electrons. Eg.  $\text{BF}_3$

**Expanded Octets** - Occur in the third row of the periodic table and beyond in elements that can form stable bonds with more than eight electrons. Eg.  $\text{XeF}_2$

**Hydrogen Bonding** - When hydrogen is attracted to electronegative atoms.

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