

## Oxidation-Reduction Titrations I: Determination of Oxalate

### Pre-lab Questions

Before beginning this experiment in the laboratory, you should be able to answer the following questions.

1. If 0.4586 g of sodium oxalate,  $\text{Na}_2\text{C}_2\text{O}_4$ , requires 33.67 mL of a  $\text{KMnO}_4$  solution to reach the end point, what is the molarity of the  $\text{KMnO}_4$  solution?  $0.4586\text{g} / 133.99\text{g/mol} = 0.003423$
2. Titration of an oxalate sample gave the following percentages: 15.33%, 15.55%, and 15.56%. Calculate the average and the standard deviation.
3. Why does the solution decolorize upon standing after the equivalence point has been reached?
4. Why is the  $\text{KMnO}_4$  solution filtered, and why should it not be stored in a rubber-stoppered bottle?
5. What volume of 0.100 M  $\text{KMnO}_4$  would be required to titrate 0.23 g of  $\text{K}_2[\text{Cu}(\text{C}_2\text{O}_4)_2] \cdot 2\text{H}_2\text{O}$ ?
6. Calculate the percent  $\text{C}_2\text{O}_4^{2-}$  in each of the following:  $\text{H}_2\text{C}_2\text{O}_4$ ,  $\text{Na}_2\text{C}_2\text{O}_4$ ,  $\text{K}_2\text{C}_2\text{O}_4$ , and  $\text{K}_3[\text{Al}(\text{C}_2\text{O}_4)_3] \cdot 3\text{H}_2\text{O}$ .

# Redox Titration Pre Lab Questions Answers

**Eugene LeMay, Jr., Herbert Beall, Karen  
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Brower, Prentice-Hall Staff**

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