

Redox Reactions Lab Reporting Worksheet

In science, reporting what has been done in a laboratory setting is incredibly important for communicating, replicating, and validating findings. However, writing scientific reports can be a little overwhelming. There are a set of agreed upon components that the scientific community requires when reporting scientific research. Answer the following questions to describe what occurred during the lab you conducted in Labster. Be sure to use complete sentences and descriptions that fully represent what you experienced. Writing a lab report is less about being correct or incorrect, than it is accurately reporting what happened and why. So, do not worry about reporting data that might seem counterintuitive or unexpected. Focus on clearly communicating what you did and what you observed.

Write your answers on a new line.

- What was the title of the lab you completed?

- *Redox Reactions: Discover how batteries work.*

1. Topic

- What was the subject you were trying to understand better in the lab?

- *Our aim was to enhance our understanding of oxidation states, elucidate periodic table trends, apply rules for determining oxidation states, describe the electron's role in redox processes, predict redox reaction directions, and balance redox reactions in acidic and alkaline conditions.*

• *We sought to deepen our comprehension of how batteries generate electricity via redox reactions. Our objective was to pinpoint optimal redox pairs to develop a more efficient and environmentally friendly battery.*

2. Background Information

- What information from the textbook and classroom is relevant for the subject you were trying to gain a better understanding of in the lab? Identify the concepts and explain how they are related to the lab topic.

- *Information that was informative from the textbook and labs:*

• *Redox reactions are characterized by alterations in the oxidation numbers of reactants, particularly in electrochemical contexts where electrons are exchanged. In aqueous environments, these processes frequently incorporate water or its constituent ions, H^+ and OH^- , alongside oxidizing and reducing agents. It was essential during our laboratory work to balance equations representing these reactions. We employed the half-reaction method, which systematically addresses the oxidation and reduction half-reactions separately, to achieve this balance.*

Redox Reactions Lab Report

Christian G. Meyer



Redox Reactions Lab Report:

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