

1. $\text{KMnO}_4 + \text{C}_2\text{H}_5\text{OH} \longrightarrow \text{Mn}^{2+} + \text{C}_2\text{H}_4\text{O} + \text{K}^+$ in acidic solution
2. $\text{KMnO}_4 + \text{C}_2\text{H}_5\text{OH} \longrightarrow \text{Mn}^{2+} + \text{C}_2\text{H}_4\text{O} + \text{K}^+$ in basic solution
3. $\text{Cr}_2\text{O}_7^{2-} + \text{Cl}^- \longrightarrow \text{Cr}^{3+} + \text{Cl}_2$ in acidic solution
4. $\text{Au}_{(\text{aq})}^{3+} + \text{I}_{(\text{aq})}^- \longrightarrow \text{Au}_{(\text{s})} + \text{I}_{2(\text{s})}$ in acidic solution
5. $\text{SO}_{3(\text{aq})}^{2-} + \text{MnO}_{4(\text{aq})}^- \longrightarrow \text{SO}_{4(\text{aq})}^{2-} + \text{Mn}_{(\text{aq})}^{2+}$ in acidic solution
6. $\text{ClO}^- + \text{Fe}(\text{OH})_3 \longrightarrow \text{Cl}^- + \text{FeO}_4^{2-}$ in basic solution

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