## **Practice Problems: Redox Reactions**

1. Determine the oxidation number of the elements in each of the following compounds:

b. 
$$N_2$$
  $N = -3$ 

f. Fe<sub>3</sub>O<sub>4</sub> 
$$3x - 8 = 0$$
  $3x = 8 + 8 | 3$   
Hint  $+ 8 + 8$   $0 = -2$   $+ 6 = + 8 | 3$ 

2. Identify the species being oxidized and reduced in each of the following reactions:

a. 
$$Cr^+ + Sn^{4+} \rightarrow Cr^{3+} + Sn^{2+}$$
  $Cv^{3+} = 0 \times idized$   
 $SN^{2+} = reduced$ 

3. Would you use an oxidizing agent or reducing agent in order for the following reactions to occur?

# **Redox Practice Problems With Answers**

**R Sandford** 

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