

## Chem 12 Electrochemistry : Quiz 2 - 120

### Part 1

1) A particular reaction has a predicted  $E^\circ$  of -2.25 V. This indicates that the reaction would :

- a) occur rapidly                      b) produce energy                      c) go to completion  
d) be non-spontaneous

2) Corrosion of a metal means that the metal :

- a) acts as a cathode                      b) undergoes oxidation  
c) undergoes reduction                      d) acts as an oxidizing agent

3) Zinc metal is prepared by electrolysis as follows:



What is the minimum voltage necessary for the electrolysis reaction to proceed ?

- a) 2.12                      b) 1.58                      c) 1.96                      d) 3.48

4) Which metal would provide the best cathodic protection for iron pipes ?

- a) Ag                      b) Sn                      c) Mg                      d) Cu

5) Which half-reaction best describes what is taking place at the anode during the electrolysis of a 1.0 M NaCl solution ?

- a)  $\text{Na}^{+} + \text{e}^{-} \rightarrow \text{Na}$                       b)  $2\text{Cl}^{-} \rightarrow \text{Cl}_2(\text{g}) + 2\text{e}^{-}$   
c)  $2\text{H}^{+}(10^{-7}\text{M}) + 2\text{e}^{-} \rightarrow \text{H}_2(\text{g})$                       d)  $\text{H}_2\text{O} \rightarrow 1/2 \text{O}_2(\text{g}) + 2\text{H}^{+}(10^{-7}\text{M}) + 2\text{e}^{-}$

6) Which one of the following converts chemical energy to electrical energy ?

- a) Industrial production of aluminum                      b) Nickel plating  
c) Discharging a "dry" cell                      d) Production of Na(s)

7) Which is the unit for the energy per coulomb of electrons ?

# Redox And Electrochemistry Quiz Answers

**Xiaolong Qi**



## **Redox And Electrochemistry Quiz Answers:**

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