

## Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

**Remember:** Limiting reagent is the reactant that will completely run out first in a chemical reaction.  
**Calculate:** Moles of product formed from a limiting reagent using stoichiometry. Percent yield = (actual yield / theoretical yield) x 100.

Chemical reactions involve the breaking and forming of chemical bonds. In a chemical reaction, the atoms of the reactants are rearranged to form the products.



The reaction of hydrogen and oxygen to form water is an example of a chemical reaction.

2H <sub>2</sub> (g)	100.00 grams
O <sub>2</sub> (g)	50.00 grams
H <sub>2</sub> O(l)	100.00 grams
H <sub>2</sub> (g)	100.00 grams

1. 100.00 g of hydrogen and 50.00 g of oxygen react to form 100.00 g of water. The limiting reagent is hydrogen. The reaction is complete when 100.00 g of hydrogen is used.



2. 100.00 g of hydrogen and 100.00 g of oxygen react to form 100.00 g of water. The limiting reagent is oxygen.

3. 100.00 g of hydrogen and 100.00 g of oxygen react to form 100.00 g of water. The limiting reagent is hydrogen.



4. 100.00 g of hydrogen and 100.00 g of oxygen react to form 100.00 g of water. The limiting reagent is oxygen. The reaction is complete when 100.00 g of oxygen is used.



# Percent Yield And Limiting Reactants Quiz

**Kaplan Test Prep**



## Percent Yield And Limiting Reactants Quiz:

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