

Name: _____

Oxidation Numbers Worksheet

Directions: Use the *Rules for Assigning Oxidation Numbers* to determine the oxidation number assigned to each element in each of the given chemical formulas.

Rules for Assigning Oxidation Numbers	Example:
1. The oxidation number of any uncombined element is 0 .	Ne, O ₂ = 0
2. The oxidation number of a monatomic ion equals the charge on the ion.	Fe ⁺³ = +3
3. The oxidation number of halogens in a compound is mostly -1 .	F ⁻¹ = -1
4. Oxygen has an oxidation number of -2 unless it's a peroxide is -1	H ₂ O ⁻² = -2 Na ₂ O ₂ ⁻² = -1
5. The oxidation number of a metal is +1 in Group 1 and +2 in Group 2.	K ⁺¹ = +1
6. Hydrogen works with +1 with nonmetals and -1 with metals.	H ⁺¹ Cl ⁻¹ = +1 Na ⁺¹ H ⁻¹ = -1
7. The sum of the oxidation numbers of all atoms in a neutral compound is 0.	
8. The sum of the oxidation numbers of all atoms in a polyatomic ion = the charge of the ion.	

1. Give oxidation numbers for the underlined atoms in these molecules:

- | | | | | | | |
|--|-----|-----|-------------------------|---------------------------------------|-----|----|
| a. <u>Cs₂O</u> | Cs: | O: | i. <u>N₂</u> | N: | | |
| b. <u>N₂O₃</u> | N: | O: | j. <u>Kr</u> | Kr: | | |
| c. <u>Na₄SiO₄</u> | Na: | Si: | O: | k. <u>H₂O</u> | H: | O: |
| d. <u>K₂Cr₂O₇</u> | K: | Cr: | O: | l. <u>FeO</u> | Fe: | O: |
| e. <u>H₂O₂</u> | H: | O: | (This is a peroxide) | m. <u>CaS</u> | Ca: | S: |
| f. <u>Al(OH)₃</u> | Al: | O: | H: | n. <u>H₂</u> | H: | |
| g. <u>HPO₃</u> | H: | P: | O: | o. <u>He</u> | He: | |
| h. <u>H₂SeO₃</u> | H: | Se: | O: | p. <u>H₂SO₄</u> | S: | H: |

2. Give the oxidation numbers for the following ions

- | | | | | | | |
|---------------------------------------|-----|---------------------------|--|---------------------------|-----|----|
| a. <u>Cu⁺¹</u> | Cu: | b. <u>Co²⁺</u> | Co: | c. <u>Cl⁻¹</u> | Cl: | |
| d. <u>IO₂⁻¹</u> | I: | O: | e. <u>SbF₆⁻¹</u> | Sb: | F: | |
| | | | | f. <u>OH⁻¹</u> | O: | H: |

Oxidation Number Problems With Answers

Chao Zhang



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Problems in Inorganic Chemistry Madan R.L., 2014 Problems in Inorganic Chemistry

Quantitative Chemical Analysis Daniel C. Harris, 2010-04-30 QCA is the bestselling textbook of choice for analytical chemistry It offers a modern portrait of the

techniques of chemical analysis backed by a wealth of real world applications This edition features new coverage of spectroscopy and statistics new pedagogy and enhanced lecturer support **Relevant Problems for Chemical Principles** Ian S. Butler, Arthur E. Grosser, 1984 **Misconceptions in Chemistry** Hans-Dieter Barke, Al Hazari, Sileshi Yitbarek, 2008-11-18 Over the last decades several researchers discovered that children pupils and even young adults develop their own understanding of how nature really works These pre concepts concerning combustion gases or conservation of mass are brought into lectures and teachers have to diagnose and to reflect on them for better instruction In addition there are school made misconceptions concerning equilibrium acid base or redox reactions which originate from inappropriate curriculum and instruction materials The primary goal of this monograph is to help teachers at universities colleges and schools to diagnose and cure the pre concepts In case of the school made misconceptions it will help to prevent them from the very beginning through reflective teaching The volume includes detailed descriptions of class room experiments and structural models to cure and to prevent these misconceptions **Mapping College Chemistry** Stephen DeMeo, 2019-03-01 This text is a chemistry problem solving resource appropriate for teachers and their students who are enrolled in high school Advanced Placement Chemistry or in a first year college General Chemistry course The book incorporates a chemistry problem solving plan one that uses an innovative graphic organizer strategy The strategy successfully evaluated with students combines problem solving processes with chemical concepts that will allow students to solve the most common and difficult problems encountered in the first year of chemistry Topical problem solving will focus on limiting reactant stoichiometry identifying types of chemical reactions equilibrium acid base equilibria and electrochemistry Why would this resource be of interest to chemistry students To be successful to get into a well known college medical school physical therapy or graduate program often requires that students get an A in your pre requisite Introductory General Chemistry course To make matters worse many college professors feel that only a few students should get A grades and therefore they give difficult exams that many students fail this is the weeding out process that every pre health student is apprehensive about To succeed in this competitive environment entails not just studying harder or longer it means re organizing textbook content so that it is meaningful to the student This is the first text of its kind to employ a reliable research based strategy that incorporates a decision based visual tool to solve chemistry textbook problems ones that can make or break a career **Electronic Structure of Materials** Natalia Chezhina, Dmitry Korolev, 2019-04-01 This book is a short survey of magnetochemistry as a promising method for revealing the electronic structure of inorganic substances particularly solid oxide materials It is supported by five chapters that describe materials with various structures and applications showing how the method of magnetic dilution with the aid of other physical methods electron spin resonance magnetization Raman and Mossbauer spectroscopy and electrical conductivity accompanied by thorough structural and quantum mechanical studies may be used for describing the states of atoms and interatomic interactions in multicomponent

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Zhihan Nan, Sheng Zhang, 2019-11-19 This study guide for the Chemistry Olympiad contains summarized concepts and examples in all areas of chemistry The chapters are arranged in a logical manner and establishes connections between concepts Undergraduate chemistry concepts are explained clearly every equation in physical chemistry is derived and justified while every organic reaction has its reaction mechanism shown and explained without assuming that readers have university level background in the subject The book also contains original Chemistry Olympiad sample problems that readers may use to test their knowledge This is a first book of its kind written by Nan Zhihan International Chemistry Olympiad IChO gold medallist and winner of the International Union of Pure and Applied Chemistry IUPAC Prize for achieving the highest score in the experimental exam and experienced Chemistry Olympiad trainer Dr Zhang Sheng who has served as head mentor of Singapore IChO team for many years It builds on the experience of both a participant and trainer to help any aspiring Chemistry Olympiad student understand the challenging concepts in chemistry

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