# **Oxidation - Reduction Reactions**

## Redox Terminology—REVIEW CHAPTER 4!!!

```
redox reaction -- electron transfer process
e.g., 2 Na + Cl<sub>2</sub> --> 2 NaCl
```

## Overall process involves two Half Reactions:

```
oxidation -- loss of electron(s)

reduction -- gain of electron(s)

e.g. Na --> Na* + e<sup>-</sup> (oxidation)

Cl<sub>2</sub> + 2 e<sup>-</sup> --> 2 Cl<sup>-</sup> (reduction)
```

## Related terms:

```
oxidizing agent = the substance that gets reduced (CI<sub>2</sub>)
reducing agent = the substance that gets oxidized (Na)
```

Oxidation and reduction always occur together so that there is no net loss or gain of electrons overall.

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# **Oxidation Reduction Guide Answers**

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