

CHAPTER 24 NUCLEAR REACTIONS AND THEIR APPLICATIONS

- 24.1 a) Chemical reactions are accompanied by relatively small changes in energy while nuclear reactions are accompanied by relatively large changes in energy.
 b) Increasing temperature increases the rate of a chemical reaction but has no effect on a nuclear reaction.
 c) Both chemical and nuclear reaction rates increase with higher reactant concentrations.
 d) If the reactant is limiting in a chemical reaction, then more reactant produces more product and the yield increases in a chemical reaction. The presence of more radioactive reagent results in more decay product, so a higher reactant concentration increases the yield in a nuclear reaction.
- 24.2 a) The percentage of sulfur atoms that are sulfur-32 is 95.02%, the same as the relative abundance of ^{32}S .
 b) The atomic mass is larger than the isotopic mass of ^{32}S . Sulfur-32 is the lightest isotope, as stated in the problem, so the other 5% of sulfur atoms are heavier than 31.972070 amu. The average mass of all the sulfur atoms will therefore be greater than the mass of a sulfur-32 atom.
- 24.3 a) She found that the intensity of emitted radiation is directly proportional to the concentration of the element in the various samples, not to the nature of the compound in which the element occurs.
 b) She found that certain uranium minerals were more radioactive than pure uranium, which implied that they contained traces of one or more as yet unknown, highly radioactive elements. Pitchblende is the principal ore of uranium.
- 24.4 Radioactive decay that produces a different element requires a change in atomic number (Z , number of protons).
- | | | | |
|-----------|-----|---|-------------------------------------|
| ${}_Z^AX$ | A | = | mass number (protons + neutrons) |
| | Z | = | number of protons (positive charge) |
| | X | = | symbol for the particle |
| | N | = | $A - Z$ (number of neutrons) |
- a) Alpha decay produces an atom of a different element, i.e., a daughter with two less protons and two less neutrons.
- $${}_Z^AX \rightarrow {}_{Z-2}^{A-4}Y + {}_2^4\text{He} \quad \text{2 fewer protons, 2 fewer neutrons}$$
- b) Beta decay produces an atom of a different element, i.e., a daughter with one more proton and one less neutron. A neutron is converted to a proton and β particle in this type of decay.
- $${}_Z^AX \rightarrow {}_{Z+1}^AY + {}_{-1}^0\beta \quad \text{1 more proton, 1 less neutron}$$
- c) Gamma decay does not produce an atom of a different element and Z and N remain unchanged.
- $${}_Z^AX^* \rightarrow {}_Z^AX + {}_0^0\gamma \quad ({}_Z^AX^* = \text{energy rich state}), \text{ no change in number of protons or neutrons.}$$
- d) Positron emission produces an atom of a different element, i.e., a daughter with one less proton and one more neutron. A proton is converted into a neutron and positron in this type of decay.
- $${}_Z^AX \rightarrow {}_{Z-1}^AY + {}_{+1}^0\beta \quad \text{1 less proton, 1 more neutron}$$
- e) Electron capture produces an atom of a different element, i.e., a daughter with one less proton and one more neutron. The net result of electron capture is the same as positron emission, but the two processes are different.
- $${}_Z^AX + {}_{-1}^0e \rightarrow {}_{Z-1}^AY \quad \text{1 less proton, 1 more neutron}$$
- A different element is produced in all cases except (c).
- 24.5 The key factor that determines the stability of a nuclide is the ratio of the number of neutrons to the number of protons, the N/Z ratio. If the N/Z ratio is either too high or not high enough, the nuclide is unstable and decays.
- | | |
|-------------------|---|
| ${}_2^3\text{He}$ | $N/Z = 1/2$ |
| ${}_2^4\text{He}$ | $N/Z = 0/2$, thus it is more unstable. |

Reinforcement Chapter 24 Nuclear Reactions

Tao Wei



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