

Appendix D

ANSWERS TO SOME PROBLEMS

1-1. (a) At standard temperature and pressure, a mole of an ideal gas contains 6.022×10^{23} molecules (Avogadro's number) and occupies 22.4 liters. Hence, the number per m^3 is $6.022 \times 10^{23}/2.24 \times 10^{-2} = 2.66 \times 10^{25} \text{ m}^{-3}$.

(b) Since $PV = NRT$, $n = N/V = P/RT$. Hence $n_1/n_0 = P_1T_0/P_0T_1$. Taking n_0 to be the density in part (a) and n_1 to be that in part (b), we have

$$n_1 = (2.66 \times 10^{25}) \frac{10^{-3}}{760} \frac{273}{(273 + 20)} = 3.30 \times 10^{19} \text{ m}^{-3}$$

Note that a diatomic gas such as H_2 will have twice as many atoms per torr as, say, He.

1-2. Consider the integral

$$I^2 = \int_{-\infty}^{\infty} e^{-x^2} dx \int_{-\infty}^{\infty} e^{-y^2} dy = \iint_{-\infty}^{\infty} e^{-(x^2+y^2)} dx dy$$

in a two-dimensional space. Transforming to cylindrical coordinates, we have

$$\begin{aligned} I^2 &= \iint e^{-r^2} r dr d\phi = 2\pi \int_0^{\infty} e^{-r^2} r dr \\ &= \pi \int_0^{\infty} e^{-r^2} d(r^2) = -\pi e^{-r^2} \Big|_0^{\infty} = \pi \end{aligned}$$

Hence,

$$I = \int_{-\infty}^{\infty} e^{-x^2} dx = \sqrt{\pi}$$

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S Ashworth



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Table of Contents Plasma Chen Solution Manual

1. Understanding the eBook Plasma Chen Solution Manual
 - The Rise of Digital Reading Plasma Chen Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Plasma Chen Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Plasma Chen Solution Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Plasma Chen Solution Manual
 - Personalized Recommendations
 - Plasma Chen Solution Manual User Reviews and Ratings
 - Plasma Chen Solution Manual and Bestseller Lists
5. Accessing Plasma Chen Solution Manual Free and Paid eBooks
 - Plasma Chen Solution Manual Public Domain eBooks
 - Plasma Chen Solution Manual eBook Subscription Services
 - Plasma Chen Solution Manual Budget-Friendly Options
6. Navigating Plasma Chen Solution Manual eBook Formats
 - ePub, PDF, MOBI, and More
 - Plasma Chen Solution Manual Compatibility with Devices
 - Plasma Chen Solution Manual Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Plasma Chen Solution Manual
 - Highlighting and Note-Taking Plasma Chen Solution Manual
 - Interactive Elements Plasma Chen Solution Manual
8. Staying Engaged with Plasma Chen Solution Manual

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Plasma Chen Solution Manual
- 9. Balancing eBooks and Physical Books Plasma Chen Solution Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Plasma Chen Solution Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Plasma Chen Solution Manual
 - Setting Reading Goals Plasma Chen Solution Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Plasma Chen Solution Manual
 - Fact-Checking eBook Content of Plasma Chen Solution Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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