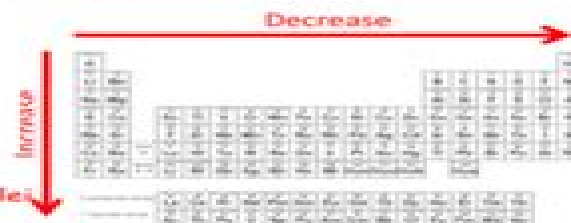


Periodic Trends Test Review

Atomic Radii (Atomic Size)

- Draw the trend lines for atomic radius on the periodic table to the right.
- What is the definition of atomic radius?
Atomic radius is one half the distance between the nuclei of two atoms (of the same element) that are bonded together.
- What happens to the size of atoms as you move down a group? **Increases**
Why? **Down a group each element has an additional (and larger) energy level to hold its electrons.**
- What happens to the size of an atom as you move left to right across a period? **Decreases**
Why? **Across a period each successive element has an additional proton which increases the positive charge in the nucleus. The positive charge attracts the negative electrons, and pulls them toward the nucleus, making the atom smaller.**
- How are the numbers of valence electrons of the Group A elements related to the group number? **The group number tells you how many valence electrons elements in that group have. For example, Group 1A elements each have 1 valence electron.**
- How is the energy level of an atom's valence electrons related to the period it is in on the periodic table? **The period number tells you the energy level that an atom's valence electron(s) is/are located. For example, Sodium (Na) is in the third period. Its valence electron is located in the 3rd energy level.**
- Put the following sets of elements in order of **increasing** atomic size (smallest to largest).

a. Aluminum, Sodium, Chlorine	Cl, Al, Na
b. Rubidium, Lithium, Potassium	Li, K, Rb
c. Tin, Iodine, Xenon	Xe, I, Sn
d. Bismuth, Phosphorous, Nitrogen	N, P, Bi
e. Beryllium, Barium, Calcium	Be, Ca, Ba
- Given any two elements within a group, is the element with the smaller or larger atomic number likely to have a larger or smaller atomic radius than the other element?
The element with the larger atomic number will have the larger atomic radius.



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