

Periodic Trends – Atomic Mass

- The **atomic mass** describes the mass of each individual element.
- **Atomic mass** increases moving **down** each group and **right** across a period.

Instructions: In the periodic table below, draw two arrows representing the direction that atomic mass **increases** across groups and periods.

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1		2											13	14	15	16	17	
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3			3	4	5	6	7	8	9	10	11	12						
4																		
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6																		
7																		

Instructions: In each question, identify the element with the **greatest** atomic mass:

- 1) Chlorine (Cl), Iodine (I), Bromine (Br), Fluorine (F) Iodine
- 2) Oxygen (O), Carbon (C), Nitrogen (N), Boron (B) Oxygen
- 3) Lithium (Li), Silicon (Si), Sulfur (Br), Xenon (Xe) Xenon

Instructions: In each question, list the elements from **least** to **greatest** atomic mass:

- 4) Magnesium (Mg), Beryllium (Be), Barium (Ba) Be, Mg, Ba
- 5) Gallium (Ga), Selenium (Se), Potassium (K) K, Ga, Se

Periodic Trends In Atomic Properties Answers

**Dr Vijay Sarda, Dr A C Handa, Dr K K
Arora**



Periodic Trends In Atomic Properties Answers:

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Table of Contents Periodic Trends In Atomic Properties Answers

1. Understanding the eBook Periodic Trends In Atomic Properties Answers
 - The Rise of Digital Reading Periodic Trends In Atomic Properties Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends In Atomic Properties Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends In Atomic Properties Answers
 - Personalized Recommendations
 - Periodic Trends In Atomic Properties Answers User Reviews and Ratings

- Periodic Trends In Atomic Properties Answers and Bestseller Lists
- 5. Accessing Periodic Trends In Atomic Properties Answers Free and Paid eBooks
 - Periodic Trends In Atomic Properties Answers Public Domain eBooks
 - Periodic Trends In Atomic Properties Answers eBook Subscription Services
 - Periodic Trends In Atomic Properties Answers Budget-Friendly Options
- 6. Navigating Periodic Trends In Atomic Properties Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends In Atomic Properties Answers Compatibility with Devices
 - Periodic Trends In Atomic Properties Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends In Atomic Properties Answers
 - Highlighting and Note-Taking Periodic Trends In Atomic Properties Answers
 - Interactive Elements Periodic Trends In Atomic Properties Answers
- 8. Staying Engaged with Periodic Trends In Atomic Properties Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends In Atomic Properties Answers
- 9. Balancing eBooks and Physical Books Periodic Trends In Atomic Properties Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends In Atomic Properties Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends In Atomic Properties Answers
 - Setting Reading Goals Periodic Trends In Atomic Properties Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends In Atomic Properties Answers
 - Fact-Checking eBook Content of Periodic Trends In Atomic Properties Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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