

Unit 2: The Periodic Table Review Sheet

U2L1: PT History

- On the present day Periodic Table, elements are arranged in order of increasing atomic number
- Groups:** vertical columns on the periodic table
 - Elements in the same group have similar chemical properties - because they have the same number of valence electrons
 - Valence electron: outermost electrons (lost # in the electron configuration)
- Periods:** horizontal rows on the periodic table

U2L2: Parts of the Periodic Table

- Elements are classified as either metals, nonmetals, or metalloids
 - Classification can be determined by location on the periodic table or by different properties

1

H

6

C

12.01

Group 14, Period 2

Valence Electrons: 4

Electron Configuration: $1s^2 2s^2 2p^2$

18

He

Group 1

Group 18

Li	Be															B	C	N	O	F	Ne
Na	Mg															Al	Si	P	S	Cl	Ar
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr				
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe				
Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn				
Fr	Ra	Ac	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn	Uut	Uuq	Uup	Uuh	Uus	Uuo				

Periodic Table Review Pt 2

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