

**Periodic Trends**

1. Explain the following trends:

- Atomic Radius
- Electronegativity
- Ionization Energy

2. Order the following elements by increasing atomic radius:

Li, Na, K, Rb, Cs, Fr

3. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

4. An element has the following properties:

- Atomic Number: 19
- Electronegativity: 0.9
- Ionization Energy: 418 kJ/mol

5. Order the following elements by increasing ionization energy:

Li, Na, K, Rb, Cs, Fr

6. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

7. Order the following elements by increasing atomic radius:

Li, Na, K, Rb, Cs, Fr

8. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

9. Order the following elements by increasing ionization energy:

Li, Na, K, Rb, Cs, Fr

10. Order the following elements by increasing electronegativity:

F, Cl, Br, I, At

Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

**Periodic Trends Test Study Guide**

**Atomic Structure and Properties of the Periodic Table**

1. What are the three subatomic particles found in the periodic table?

- Protons: positively charged, found in the nucleus.
- Neutrons: neutral, found in the nucleus.
- Electrons: negatively charged, found in the electron cloud.

2. Label the following parts of the periodic table: s-block, p-block, d-block, f-block, transition metals, main group elements, noble gases, halogens, alkali metals, lanthanides, actinides.

3. Complete the following table:

Block	Groups	Periods	Elements
s-block	1, 2	1-7	Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Ra
p-block	13-18	2-7	B, C, N, O, F, Ne, Al, Si, P, S, Cl, Ar, As, Se, Br, Kr, Sb, Te, I, Xe, Bi, Po, At, Rn
d-block	3-10	4-7	Sc, Ti, V, Cr, Mn, Fe, Co, Ni, Cu, Zn, Ga, Ge, As, Se, Br, Kr, Rb, Sr, Y, Zr, Nb, Mo, Tc, Ru, Rh, Pd, Ag, Cd, In, Sn, Sb, Te, I, Xe, Ba, La, Ce, Pr, Nd, Pm, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb, Lu, Hf, Ta, W, Re, Os, Ir, Pt, Au, Hg, Tl, Pb, Bi, Po, At, Rn, Ac, Th, Pa, U, Np, Pu, Am, Cm, Bk, Cf, Es, Fm, Md, No, Lr
f-block	14-15	6-7	Ce, Pr, Nd, Pm, Sm, Eu, Gd, Tb, Dy, Ho, Er, Tm, Yb, Lu, Th, Pa, U, Np, Pu, Am, Cm, Bk, Cf, Es, Fm, Md, No, Lr

4. Explain the periodic trends of atomic radius, electronegativity, ionization energy, and electron affinity.

5. Complete the following table:

Element	Atomic Number	Electronegativity	Ionization Energy (kJ/mol)	Electron Affinity (kJ/mol)
Hydrogen	1	2.20	1312	-72
Helium	2	2.20	2372	0
Lithium	3	0.98	520	-60
Boron	5	2.04	801	-27
Carbon	6	2.55	1086	-122
Nitrogen	7	3.04	1402	-70
Oxygen	8	3.44	1314	-141
Fluorine	9	3.98	1681	-328
Neon	10	4.79	2081	0
Sodium	11	0.93	496	-48
Magnesium	12	1.31	738	-74
Aluminum	13	1.61	578	-42
Silicon	14	1.90	786	-134
Phosphorus	15	2.19	1012	-72
Sulfur	16	2.58	1000	-200
Chlorine	17	3.16	1251	-349
Argon	18	3.99	1521	0
Potassium	19	0.82	419	-48
Calcium	20	1.00	590	-29
Scandium	21	1.36	633	-29
Titanium	22	1.54	658	-38
Vanadium	23	1.63	673	-51
Chromium	24	1.66	686	-64
Manganese	25	1.55	717	-74
Iron	26	1.83	762	-16
Cobalt	27	1.88	760	-32
Nickel	28	1.91	737	-112
Copper	29	1.90	746	-119
Zinc	30	1.65	738	-113
Gallium	31	1.81	579	-29
Germanium	32	2.01	609	-119
Arsenic	33	2.19	631	-78
Selenium	34	2.55	625	-200
Bromine	35	2.96	625	-325
Krypton	36	3.00	648	0
Rubidium	37	0.82	403	-47
Strontium	38	1.00	549	-29
Yttrium	39	1.22	579	-29
Zirconium	40	1.39	580	-38
Niobium	41	1.46	578	-51
Molybdenum	42	1.68	583	-64
Technetium	43	1.70	579	-74
Ruthenium	44	1.88	581	-16
Rhodium	45	1.90	579	-32
Palladium	46	1.95	566	-112
Silver	47	1.93	569	-119
Cadmium	48	1.69	566	-113
Indium	49	1.85	558	-29
Sn	50	2.02	579	-119
Sb	51	2.20	579	-78
Te	52	2.54	579	-200
I	53	2.66	548	-325
Xe	54	2.60	569	0
Cs	55	0.79	376	-47
Ba	56	0.89	503	-29
La	57	1.10	533	

# Periodic Table And Chemistry Study Guide Answers

**SB Merriam**



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