Chemistry--Unit 10: Chemical Periodicity Practice Problems

- I. Classification of the elements
- Use the periodic table to write the electron configuration for silicon.

Is 2s 2p 3s 3p or abbreviated form : [Ne] 3s 3p may use this form now!!!

2) Use the periodic table to write the electron configuration for iodine.

1s22s2p63s23p64s23d104p65s24d105p5 or [Kr]5s24d105p5

3) Which family of elements is characterized by an s²p³ configuration?

Nitrogen Family, or Group 15

- 4) Name the element that matches the following description:
 - a. one that has 5 outer electrons on the third period of the periodic table P, phosphorus
 - b. one with a 4s²4n⁵ electron configuration Br, bromine
 - c. the 4th period Group 16 element Se, selenium
- 5) Identify the elements that have the following outer electron configurations?
 - a. $2s^22p^4$ O, oxygen
 - b. $4s^2$ Ca. Calcium c. $4s^23d^{89}$ Zn, zinc
- 6) What is the common characteristic of the electron configurations of the elements Ne and Ar? In which Group would you find them?

full outermost energy level, Group 18, the noble or inert gases

7) Why would you expect lithium (Li) and sulfur (S) to have different chemical and physical properties?

they are in different groups on the periodic table (Li group 1, S group 16)

8) What characterizes silver (Ag) and iron (Fe) as transition metals?

Their electron configurations both end in the d sublevel (last electrons added are to a d subjevel).

- II. Periodic Trends
- 9) Explain why a magnesium atom is smaller than atoms of both sodium and calcium.

Na has less +/- attraction; Ca has another energy level

10) Predict the size of the astatine (At) atom compared to that of tellurium (Te).

At would be larger than Te due to the extra energy level.

Periodic Table 6 Practice Problems Answer Key

Raffaela Di Napoli

Periodic Table 6 Practice Problems Answer Key:

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material focusing on your specific problem areas The course review includes the information you need to know when taking the exam Be sure to take the remaining practice tests to further test yourself and become familiar with the format of the SAT II Chemistry Subject Test When Should I Start Studying It is never too early to start studying for the SAT II Chemistry test The earlier you begin the more time you will have to sharpen your skills Do not procrastinate Cramming is not an effective way to study since it does not allow you the time needed to learn the test material The sooner you learn the format of the exam the more comfortable you will be when you take the exam FORMAT OF THE SAT II CHEMISTRY The SAT II Chemistry is a one hour exam consisting of 85 multiple choice questions. The first part of the exam consists of classification questions This question type presents a list of statements or questions that you must match up with a group of choices lettered A through E Each choice may be used once more than once or not at all The exam then shifts to relationship analysis questions which you will answer in a specially numbered section of your answer sheet You will have to determine if each of two statements is true or false and if the second statement is a correct explanation of the first The last section is composed strictly of multiple choice questions with choices lettered A through E Material Tested The following chart summarizes the distribution of topics covered on the SAT II Chemistry Subject Test Topic Percentage Number of Questions Atomic however do not review too much at any one time Concentrate on one problem area at a time by reviewing the questions and explanations and by studying our review until you are confident you completely understand the material Keep track of your scores By doing so you will be able to gauge your progress and discover general weaknesses in particular sections You should carefully study the reviews that cover your areas of difficulty as this will build your skills in those areas TEST TAKING TIPS Although you may be unfamiliar with standardized tests such as the SAT II Chemistry Subject Test there are many ways to acquaint yourself with this type of examination and help alleviate your test taking anxieties Become comfortable with the format of the exam When you are practicing to take the SAT II Chemistry Subject Test simulate the conditions under which you will be taking the actual test Stay calm and pace yourself After simulating the test only a couple of times you will boost your chances of doing well and you will be able to sit down for the actual exam with much more confidence Know the directions and format for each section of the test Familiarizing yourself with the directions and format of the exam will not only save you time but will also ensure that you are familiar enough with the SAT II Chemistry Subject Test to avoid nervousness and the mistakes caused by being nervous Do your scratchwork in the margins of the test booklet You will not be given scrap paper during the exam and you may not perform scratchwork on your answer sheet Space is provided in your test booklet to do any necessary work or draw diagrams If you are unsure of an answer guess However if you do guess guess wisely Use the process of elimination by going through each answer to a guestion and ruling out as many of the answer choices as possible By eliminating three answer choices you give yourself a fifty fifty chance of answering correctly since there will only be two choices left from which to make your guess Mark your answers in the appropriate spaces on the

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