

Graphing Periodic Trends Activity

Name: _____ Period: _____

Use the information in these tables to complete the graph as described below

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
H	31	1312	2.1
Li	128	520	1.0
Na	166	496	0.9
K	203	410	0.8
Rb	220	403	0.8
Cs	244	376	0.7

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
Na	166	496	0.9
Mg	141	738	1.2
Al	121	578	1.5
Si	111	787	1.8
P	107	1012	2.1
S	105	1000	2.5
Cl	102	1251	3.0
Ar	106	1521	---

1. Using colored pencils or pen, list under the symbol (in this order!) the Atomic Radius, First Ionization Energy, and Electronegativity.
2. Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
3. Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
4. Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)

H	Across a period (→) atomic radius tends to _____. Across a period (→) first ionization energy tends to _____. Across a period (→) electronegativity tends to _____.						
Li							
Na 166 496 0.9	Mg	Al	Si	P	S	Cl	Ar
K	Down a Group (↓) atomic radius tends to _____. Down a Group (↓) first ionization energy tends to _____. Down a Group (↓) electronegativity tends to _____.						
Rb							
Cs							

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Hussin A.Rothana



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Table of Contents Periodic Trends Packet Answer Key Tables And Graphs

1. Understanding the eBook Periodic Trends Packet Answer Key Tables And Graphs
 - The Rise of Digital Reading Periodic Trends Packet Answer Key Tables And Graphs
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Packet Answer Key Tables And Graphs
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Packet Answer Key Tables And Graphs
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Packet Answer Key Tables And Graphs
 - Personalized Recommendations
 - Periodic Trends Packet Answer Key Tables And Graphs User Reviews and Ratings
 - Periodic Trends Packet Answer Key Tables And Graphs and Bestseller Lists
5. Accessing Periodic Trends Packet Answer Key Tables And Graphs Free and Paid eBooks
 - Periodic Trends Packet Answer Key Tables And Graphs Public Domain eBooks
 - Periodic Trends Packet Answer Key Tables And Graphs eBook Subscription Services
 - Periodic Trends Packet Answer Key Tables And Graphs Budget-Friendly Options

6. Navigating Periodic Trends Packet Answer Key Tables And Graphs eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Packet Answer Key Tables And Graphs Compatibility with Devices
 - Periodic Trends Packet Answer Key Tables And Graphs Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Packet Answer Key Tables And Graphs
 - Highlighting and Note-Taking Periodic Trends Packet Answer Key Tables And Graphs
 - Interactive Elements Periodic Trends Packet Answer Key Tables And Graphs
8. Staying Engaged with Periodic Trends Packet Answer Key Tables And Graphs
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Packet Answer Key Tables And Graphs
9. Balancing eBooks and Physical Books Periodic Trends Packet Answer Key Tables And Graphs
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Packet Answer Key Tables And Graphs
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Packet Answer Key Tables And Graphs
 - Setting Reading Goals Periodic Trends Packet Answer Key Tables And Graphs
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Packet Answer Key Tables And Graphs
 - Fact-Checking eBook Content of Periodic Trends Packet Answer Key Tables And Graphs
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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