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Service Manual

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mole H_2S x 18g H_2O i mole H_2O 149 4g H_2O is your theoretical yield so percent yield $\frac{137 \text{ g } \text{H}_2\text{O}}{149 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$ problems 6 8 are similar to problem 5

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ 4 Al + 3 $\text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al x $\frac{1 \text{ mol Al}}{27 \text{ g/mol}}$ x $\frac{2 \text{ mol } \text{Al}_2\text{O}_3}{4 \text{ mol Al}}$ x 101.964 g Al_2O_3 302.3 g Al_2O_3 26.982 g Al 4 mol Al 1 mol Al_2O_3 c determine the percent yield $\frac{\text{yield actual}}{\text{yield theoretical}} \times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$ theoretical 302.3 g

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web 7 Na + 2 C + 2 O₄ is the limiting reactant percent yield 86.6% 8 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting reagent 4 33 g of H₂ are left over 13

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