

## Finding the Percent Composition

- **Percent Composition** - The percentage (by mass) of each element in a compound.

### **Example 1:**

Find the percent composition of red candies in a bag of 200 candies, if there are 26 red, and 174 blue candies.

- Total Candies: 200
- Red Candies: 26
- Blue Candies: 174

$$\frac{26 \text{ red}}{200 \text{ total}} \times 100 = 13.0 \% \text{ red}$$

Formula:

$$\frac{\text{Part}}{\text{Total}} \times 100$$

$$\frac{174 \text{ blue}}{200 \text{ total}} \times 100 = 87.0 \% \text{ blue}$$

### **Example 2:**

Find the percent composition of  $\text{MgCl}_2$ .

- Total mass  $\text{MgCl}_2$ : 95.21 g
- Mass of Mg: 24.31 g
- Mass of  $\text{Cl}_2$ : 70.91 g

$$\frac{24.31 \text{ g}}{95.21 \text{ g}} \times 100 = 25.53 \% \text{ Mg}$$

$$\frac{70.91 \text{ g}}{95.21 \text{ g}} \times 100 = 74.47 \% \text{ Cl}_2$$

### **Example 3:**

Find the percent composition of  $\text{H}_2\text{SO}_4$ .

- Total mass  $\text{H}_2\text{SO}_4$ : 98.08 g
- Mass of  $\text{H}_2$ : 2.02 g
- Mass of S: 32.07 g
- Mass of  $\text{O}_4$ : 64.00 g

$$\frac{2.02 \text{ g}}{98.08 \text{ g}} \times 100 = 2.06 \% \text{ H}_2$$

$$\frac{32.07 \text{ g}}{98.08 \text{ g}} \times 100 = 32.70 \% \text{ S}$$

$$\frac{64.00 \text{ g}}{98.08 \text{ g}} \times 100 = 65.25 \% \text{ O}_4$$

# Percent Composition Answers

**Michael Hoelscher**



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