

THERMOCHEMISTRY ASSIGNMENT

1. Thermodynamic stability of graphite is defined as the heat required for conversion of graphite into diamond. The enthalpy change for this reaction is, 1.9 kJ mol^{-1} . Sublimation enthalpy of graphite is 713 kJ mol^{-1} . Calculate the enthalpy of melting of diamond. (Graphite is more stable than diamond)

If the above reaction is reversed, what will occur with release of energy and increase the temperature in the hot water, the graphite will be melting produced. As a result the hot water will be cooled down and the graphite will get into liquid state.

2. Thermodynamic heat capacity of diamond is $12.485 \text{ J K}^{-1} \text{ mol}^{-1}$. Heat conductivity is used to raise the temperature of an 8.25 g block of diamond from 20.0°C to 100.0°C .

(a) Calculate

$$(i) \text{ Heat capacity of diamond } = 12.485 \times 8.25 = 102.82 \text{ J K}^{-1}$$

(ii) $\Delta T = 100.0 - 20.0 = 80.0^\circ\text{C}$

(iii) $Q = C_p \Delta T$

3. Thermodynamic heat capacity of graphite is $8.53 \text{ J K}^{-1} \text{ mol}^{-1}$. Minimum thermodynamic temperature for which the temperature of 500 g block of graphite is 100°C .

(a) $n = 500 \text{ g}$

$$(i) \text{ Heat capacity of } 500 \text{ g of graphite} = 8.53 \times 500 = 4265 \text{ J K}^{-1}$$

(ii) $\Delta T = 100.0 - 0.0 = 100.0^\circ\text{C}$

(iii) $Q = C_p \Delta T$

$$Q = 4265 \times 100 = 426500 \text{ J}$$

- (b) what the temperature of 500 g of graphite from 20.0°C to 100.0°C .

$$Q = C_p \Delta T \Rightarrow 426500 = 4265 \times \Delta T$$

$$\Delta T = \frac{426500}{4265} = 100.0^\circ\text{C}$$

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$$Q = C_p \Delta T \Rightarrow 426500 = 4265 \times \Delta T$$

$$\Delta T = \frac{426500}{4265} = 100.0^\circ\text{C}$$

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4. To raise 75.0 g of water the temperature of 80.0°C to 100.0°C it is heated by 158 kJ . What is thermodynamic heat capacity of water?

$$Q = C_p \Delta T$$

$$158000 = C_p \times 20.0$$

$$C_p = \frac{158000}{20.0} = 7900 \text{ J K}^{-1}$$

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5. Acetylene is usually released 1.54 g of water at a temperature of 25.0°C with gas production pressure. After at 25.0°C is added to the water, the final temperature is 25.0°C . Calculate the heat of solution (the enthalpy change accompanying the dissolution of the gas of gaseous acetylene).

Reaction Rates And Thermochemistry Answers

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