Equilibrium Practice Problems

 Write the equilibrium expression for each of the following reactions: $N_{2,(g)} + 3 H_{2,(g)} \leftrightarrow 2 NH_{3,(g)}$

$$K = \frac{[NH_3]^2}{[N_2][H_2]^3}$$

$$I_{2(s)} + CI_{2(g)} \leftrightarrow 2ICI_{(g)}$$

$$K = \underline{[ICI]^2}$$

$$NO_{2(q)} \leftrightarrow NO_{(q)} + V_2 O_{2(q)}$$

$$K = [NO][O_2]^{1/2}$$

 $[NO_2]$

The dissociation of acetic acid, CH₃COOH, has an equilibrium constant at 25°C of 1.8 x 10°5. The reaction is

$$CH_3COOH_{(aq)} \leftrightarrow CH_3COO^-_{(aq)} + H^+_{(aq)}$$

If the equilibrium concentration of CH₂COOH is 0.46 moles in 0.500 L of water and that of CH₃COO is 8.1 x 10⁻³ moles in the same 0.500 L, calculate [H+] for the reaction.

$$K = [CH_3COO+][H+]$$

 $K = [CH_3COO^*][H^*]$ Plug in known values and solve.

$$1.8 \times 10^{-5} = [8.1 \times 10^{-3} \text{ moles/0.500 L}] [H^+]$$

[0.46 moles/0.500 L]

$$[H^+] = 1.0 \times 10^{-3} M$$

Indicate the effect of a catalyst, pressure, temperature and concentration on each of the following on:

		catalyst	pressure	temperature	concentration
a.		presence increase	direct relationship	direct relationship	direct relationship
ь.	posn	nothing	based on moles	based on enthalpy of rxn	inverse relationship

Reaction Rates And Equilibrium Practice Problems

Barron's Educational Series, Mark Kernion, Joseph A. Mascetta

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