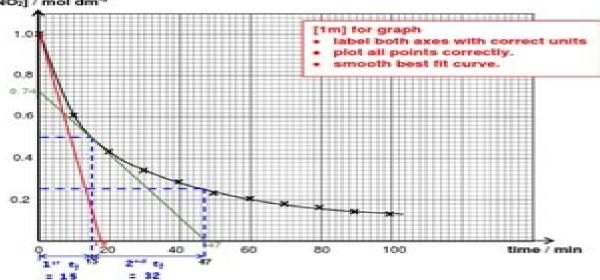
Topic 6: Reaction Kinetics (Answers)

1 (a) [NO₂] / mol dm⁻³



[1m] Indicate and label values of two half lives clearly on the graph using construction lines

Since half-lives are not constant, the reaction is not first order. [1m]

(b) (i) Initial rate = 1.0 = 0.0556 mol dm⁻³ min⁻¹ accept range of 0.0500 to 0.0625

[1m] tangent line drawn at t = 0 + value in 3st + units

(iii) rate when [NO₂] = 0.5 mol dm⁻³ = 0.74 | -0.0157 mol dm⁻³ min⁻¹

[1m] tangent line drawn at [NO₂] = 0.5 mol dm⁻¹ + value in 3sf + units

(c) [NO₂] / mol dm⁻³ rate / mol dm⁻³ min⁻¹
1.00 0.0556
0.50 0.0145

When [NO₂] increases to 2 times its original value, the rate increases to ~ 4 times its original value.

Therefore, the order of reaction with respect to [NO₂] is 2. [1m]

Reaction Energy Reaction Kinetics Answers

Roman Wölfel

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