Equilibrium Practice Problems

1. Write the equilibrium expression for each of the following reactions:

$$N_{2 (g)} + 3 H_{2 (g)} \leftrightarrow 2 NH_{3 (g)}$$

$$K = \frac{[NH_3]^2}{[N_2][H_2]^3}$$

$$I_{2 (s)} + CI_{2 (g)} \leftrightarrow 2 ICI_{(g)}$$

$$K = \underline{\prod_{[Cl_2]}}$$

$$NO_{2(g)} \leftrightarrow NO_{(g)} + 1/2 O_{2(g)}$$

$$K = [NO][O_2]^{1/2}$$

 $[NO_2]$

 The dissociation of acetic acid, CH₃COOH, has an equilibrium constant at 25°C of 1.8 x 10⁻⁵. The reaction is

If the equilibrium concentration of CH_3COOH is 0.46 moles in 0.500 L of water and that of CH_3COO' is 8.1 x 10^{-3} moles in the same 0.500 L, calculate [H $^+$] for the reaction.

$$K = [CH_3COO^*][H^*]$$

$$[CH_3COOH]$$

Plug in known values and solve.

$$1.8 \times 10^{-5} = [8.1 \times 10^{-3} \text{ moles/0.500 L}] [H^*]$$

[0.46 moles/0.500 L]

$$[H^+] = 1.0 \times 10^{-3} M$$

Indicate the effect of a catalyst, pressure, temperature and concentration on each of the following on:

		catalyst	pressure	temperature	concentration
a.	speed	presence increase	direct relationship	direct relationship	direct relationship
ь.	pos"	nothing	based on moles	based on enthalpy of rxn	inverse relationship

Reaction Rates And Equilibrium Practice Problems Answers

G. F. Azzone

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