

Periodicity - Electronegativity

Abstract

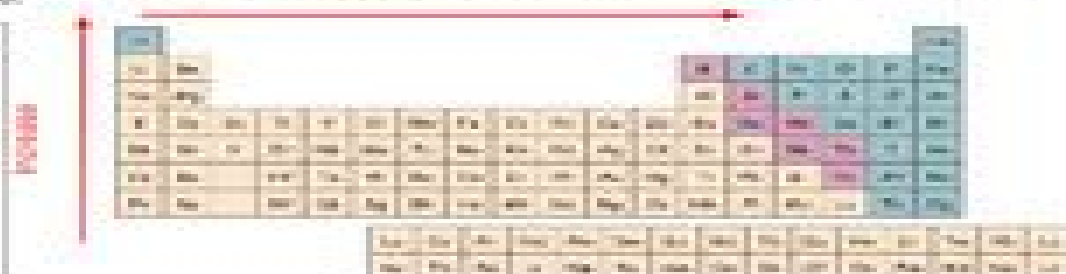
1. Fill in the blanks below, and then draw arrows on to the periodic table indicating the directions of the electronegativity trend.

[illegible]

Abstract

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2) List the elements below in order of increasing electronegativity. (5 points)

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Figure 1

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Figure 1

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Figure 1

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3. Are the elements below in order of decreasing electronegativity? If yes, write "yes". If not, write them in the correct order.

Figure 1

1998

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2000 2001 2002 2003

1998

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4. Effectively, no additional testing will be necessary for the following reasons:

Generally, electronegativity increases as you move up in a group. This is because the atomic number decreases, and as a result, there is less distance between the nucleus and valence electrons. In other words, there is a smaller atomic radius.

5. **Effectively** **manages** **and/or** **allocates** **responsibility** **effectively** **across** **network**

As you move from left to right, across a period, the nuclear charge of elements increases faster than electron shielding. This results in an increased attraction to electrons.

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