

Unit 11: Chapter 7 - Chemical Reactions Review Answer Key

Define:

1. **Reactant:** *A substance that undergoes change in a chemical reaction.*
2. **Product:** *New substances formed as a result of a chemical reaction.*
3. **Catalyst:** *A substance that affects the rate of a chemical reaction without being used up in the reaction.*
4. **Coefficient:** *Numbers that appear before a formula in a chemical equation to show the relative proportions of each reactant and product.*
5. **Subscript:** *A number written lower in a chemical formula indicating the amount of atoms present of an element (this number is not to be changed once written in chemical formula).*
6. **Law of Conservation of Mass:** *Matter cannot be created nor destroyed; only changed (atoms can be rearranged).*
7. **Synthesis reaction (include an example):** *A chemical reaction in which two or more substances react to form a single substance.* $(A + B \rightarrow AB)$
8. **Decomposition reaction (include an example):** *A chemical reaction in which a compound breaks down into two or more simpler substances.* $(AB \rightarrow A + B)$
9. **Single replacement (include an example):** *A chemical reaction in which one element takes the place of another element in a compound.* $(AB + C \rightarrow AC + B)$
10. **Double replacement (include an example):** *A chemical reaction in which two compounds exchange positive ions and form two new compounds.* $(AB + CD \rightarrow AD + CB)$
11. **Combustion (include an example):** *A chemical reaction in which a substance reacts rapidly with oxygen, often producing heat and light.* $(C_2H_2 + O_2 \rightarrow CO_2 + H_2O)$
12. **Endothermic reaction:** *A chemical reaction that absorbs energy from its surroundings.*
13. **Exothermic reaction:** *A chemical reaction that releases energy to its surroundings.*
14. **Chemical equation (include an example):** *A representation of a chemical reaction in which the reactants and products are expressed as formulas.*
15. **A balanced equation** is an equation in which the number of atoms on the reactants' side equals the number of atoms on the products' side of the chemical equation.
16. **Equilibrium:** *A state in which the forward and reverse paths of a physical or chemical change take place at the same rate.*

Reactions Review Answer

Wolfgang Guggemos

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