

Pre-Lab Information

Purpose	Explore the rate of a chemical reaction at different temperatures and with different reactant particle sizes.
Time	Approximately 45 minutes
Question	What are the effects of temperature and a reactant's particle size on reaction rate?
Hypothesis 1	If you increase the temperature of a reaction, then the reaction rate will increase because particles experience more collisions at higher temperatures.
Hypothesis 2	If you decrease the particle size of a reactant, then the reaction rate will increase because more of the reactants' surface area is exposed allowing more particles to make contact with each other.
Variables	Independent variables: particle size, temperature Dependent variable: reaction rate
Summary	The reaction between effervescent tablets and water will produce CO_2 gas, which you can observe in the form of bubbles; when the bubbling stops, the reaction is complete. You will measure the reaction times and rates for tablets dissolved at four temperatures and broken into three different particle sizes.

Safety

- Always wear a lab coat and safety goggles when performing an experiment.
- Behavior in the lab needs to be purposeful. Use caution when using the hot plate, as it can cause burns.
- Check glassware such as graduated cylinders and beakers for cracks and chips prior to use.
- Use the right gear, such as chemically resistant gloves and oven mitts, when performing the experiment.
- Report all accidents—no matter how big or small—to your teacher.

Procedure

Step 1: Gather Materials

- 250 mL graduated cylinder
- Thermometer
- Water
- Timer
- Four 250 mL beakers
- Seven 1,000 mg effervescent tablets
- Two pieces of filter paper
- 600 mL beaker
- Ice
- Hot plate

Reaction Rates Chemical Study Guide

James E. Brady, Fred Senese



Reaction Rates Chemical Study Guide:

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Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students Organic Chemistry,

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