

Lesson 19.4 Reaction rates & equilibrium

activity 1 : Answer the True or False for each question and explain the false answers.

- all reactions happen at the same rate. _____
- Reaction rate refers to how quickly the reaction takes place. _____
- Reaction rate is always given by measuring how quickly products are formed. _____
- The collision model explains that reactions occur randomly. _____

activity 2 : Identify the factors that can affect the rate of a chemical reaction.

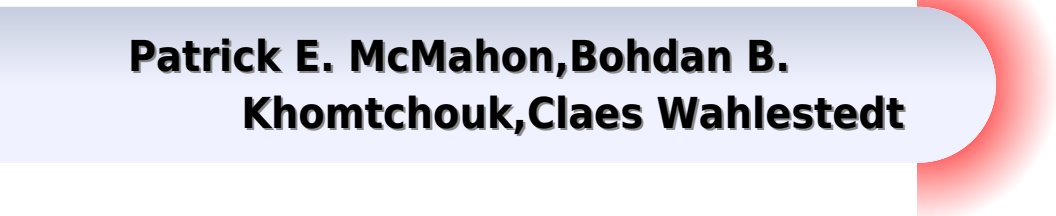
- | | |
|----------|----------|
| a. _____ | b. _____ |
| c. _____ | d. _____ |
| e. _____ | |

activity 3 : Fill in the blanks.

There are 3 _____ that control the _____ of the reaction.
 Increased in temperature _____ the speed at which particles _____ (speeding up) particles
 also increases their chance of _____ (colliding) at higher temperatures.
 Particles move _____ and _____ more frequently.
 _____ describes the number of particles of a substance in a given volume. It is often
 expressed as moles of a substance per litre (_____). Particles are more likely to _____ with
 the another at higher concentrations. Therefore there are _____ chance for collisions/increased
 reacting particles, the overall the rate of _____ in the same chemical reaction.
 For reactions involving _____ volume and pressure are important factors. Increasing the volume
 of a substance increases the _____ of gases. This means they are more likely to collide with
 each other. This increases the _____ It also increases the _____ increase
 reacting particles collide with each other more frequently.
 Increased sugar concentration speeds because individual going off sugar has much greater total
 _____ than sugar cubes.

Reaction Rates And Equilibrium Guided Answers

**Patrick E. McMahon, Bohdan B.
Khomtchouk, Claes Wahlestedt**



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