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POGIL Ionic Bonds

What are the structural units that make up ionic compounds and how are they named?

Why?

Chemical Bonding is the force that holds two atoms together. The **properties** of substances **depend** mainly on the **types of bonds** they have. Some properties of Ionic Compounds (FILL IN THE BLANKS)

Crystalline solid.
Very **High** melting point.
Soluble in H_2O . (insoluble or soluble)
Insoluble in nonpolar solvents. (insoluble or soluble)
Conduct electricity in aqueous solutions.

We will explore how ionic compounds form and how to name them later on, but we must first learn what an ion is, the different types of ions, and the importance of valence electrons in bonding.

MODEL 1: THOSE VALENCE ELECTRONS

In **IONIC BONDING** the **valence electrons** are completely **transferred** from one atom to the other atom. Ionic bonds occur between metals and nonmetals when there is a large difference in electronegativity.

What are Valence Electrons? Valence Electrons are the **s** and **p** electrons in the highest occupied energy level and are the electrons that form bonds. Bonds form in order to satisfy the **Octet Rule**. The Octet Rule states that in forming compounds, atoms tend to gain, share, or lose valence electrons to achieve a noble gas configuration, or 8 electrons in the outer level. Having 8 valence electrons in the outer level fills the valence shell, thus making the atom stable.

When a metallic element and a nonmetallic element react with each other the result is usually an **ionic compound** possessing an **ionic bond** with no charge.

Figure 1:



Key Content Questions:

1. Where are metals found in the periodic table?

On the left side of the
metalloid

Pogil Ionic Bonds Answers

**Michael P. Garoutte, Ashley B.
Mahoney**



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