

Calculating pH

How will it be used for the concentration differences here?

Why?

As buffers and other substances control pH in organisms, it is important to understand the ability of buffers to resist change. This property is called **buffering capacity** and is usually expressed as millimoles per litre of buffer that will neutralise added acid or base. This value is usually given as mmol L^{-1} of H^+ or OH^- ions. Below is a table that shows how to calculate buffer capacity for a monobasic buffer and for a dibasic buffer.

Table 1 – Ion Concentrations for Buffers and Bases

Species	Substance	Formula, Molar or Molarities	$[\text{H}^+]$ (M)	$[\text{OH}^-]$ (M)	$[\text{H}^+] + [\text{OH}^-]$ (M)
1	0.1 M 0.1 M HCl/NaCl	Acidic	0.10000×10^0	1.000×10^{-14}	0.10000
2	0.001 M 0.1 M HCl/NaCl	Acidic	0.00100×10^0	1.000×10^{-14}	0.00100
3	0.00001 M 0.1 M HCl/NaCl	Acidic	0.00001×10^0	1.000×10^{-14}	0.00001
4	0.000001 M 0.1 M HCl/NaCl	Acidic	0.000001×10^0	1.000×10^{-14}	0.000001
5	0.0000001 M 0.1 M HCl/NaCl	Neutral	1.000×10^{-7}	1.000×10^{-7}	0.0000002
6	0.00000001 M 0.1 M HCl/NaCl	Basic	1.000×10^{-8}	1.000×10^{-6}	0.00000099
7	0.000000001 M 0.1 M HCl/NaCl	Basic	1.000×10^{-9}	1.000×10^{-5}	0.000000999
8	0.0000000001 M 0.1 M HCl/NaCl	Basic	1.000×10^{-10}	1.000×10^{-4}	0.0000009999
9	0.00000000001 M 0.1 M HCl/NaCl	Basic	1.000×10^{-11}	1.000×10^{-3}	0.00000099999
10	0.000000000001 M 0.1 M HCl/NaCl	Basic	1.000×10^{-12}	1.000×10^{-2}	0.000000999999

1. What does the number $[\text{H}^+]$ in the table represent?

It represents the concentration of H^+ ions.

2. In the table 1, which one has a higher concentration of hydroxide ions ($[\text{OH}^-]$)?

0.00000000001 M

3. How did you calculate the concentration of hydroxide ions and what did you use? (Use the table 1 as a guide.)

From the $[\text{H}^+]$ concentration and the relationship between $[\text{H}^+]$ and $[\text{OH}^-]$ concentration in water: $[\text{H}^+] [\text{OH}^-] = 1.0 \times 10^{-14}$.

4. How did you calculate the concentration of hydroxide ions and what did you use? (Use the table 1 as a guide.)

From the $[\text{H}^+]$ concentration and the relationship between $[\text{H}^+]$ and $[\text{OH}^-]$ concentration in water: $[\text{H}^+] [\text{OH}^-] = 1.0 \times 10^{-14}$.

Pogil Activities Calculating Ph Answers

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