

### Periodic Trends Practice

1. Identify each element as a metal, metalloid, or nonmetal

germanium \_\_\_\_\_ **metalloid** \_\_\_\_\_

zinc \_\_\_\_\_ **metal** \_\_\_\_\_

phosphorous \_\_\_\_\_ **nonmetal** \_\_\_\_\_

lithium \_\_\_\_\_ **metal** \_\_\_\_\_

2. Give two examples of elements for each category

noble gases \_\_\_\_\_ **He, Ne, Ar, Kr, Xe, Rn** \_\_\_\_\_

halogens \_\_\_\_\_ **F, Cl, Br, I, At** \_\_\_\_\_

alkali metals \_\_\_\_\_ **Li, Na, K, Rb, Cs, Fr** \_\_\_\_\_

alkaline earth metals \_\_\_\_\_ **Be, Mg, Ca, Sr, Ba, Ra** \_\_\_\_\_

transition metals \_\_\_\_\_ **Any elements from 3B-28** \_\_\_\_\_

3. What trend in atomic radius do you see as you go down a group/family on the periodic table? What causes this trend? **Increases going down due to added energy levels. The inside energy levels are filled up so all remaining electrons must go into open spaces further away from the nucleus.**

4. What trend in atomic radius do you see as you go across a period/row on the periodic table? **Decreases going across due to electrons being added into the same valence shell. Protons would be added to the nucleus, making the positive charge stronger and electrons placed in the same space would be pulled in closer by the nucleus.**

5. Circle the atom in each pair that has the largest atomic radius.

A. **A** B

B. **S** O

C. **Br** Cl

D. **Na** Al

E. **O** F

F. Mg **Ca**

6. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

\_\_\_\_\_ **O** \_\_\_\_\_ < \_\_\_\_\_ **C** \_\_\_\_\_ < \_\_\_\_\_ **Al** \_\_\_\_\_ < \_\_\_\_\_ **K** \_\_\_\_\_

7. Define ionization energy. **The amount of energy required to remove an electron from a neutral atom**

8. Is it easier to form a positive ion with an element that has a high ionization energy or an element that has a low ionization energy? Explain. **It would be easier to form a + ion with a low ionization energy because less energy would need to be used to remove the electron.**

9. What trend in ionization energy do you see as you go down a group/family on the periodic table? **Going down a group the ionization energy decreases. The valence electrons are further away and easier to take away.**

10. What trend in ionization energy do you see as you go across a period/row on the periodic table? What causes this trend? **Going across the period, ionization energy increases. The atoms are smaller, so the nucleus is pulling electrons in closer making them harder to remove. The atoms are also increasing their number of valence electrons. As they get closer to having a complete set of 8 valence electrons they do not want to lose electrons any more, they want to gain them instead.**

11. Circle the atom in each pair that has the greater ionization energy

A. Li **Be**

B. **Ca** Ba

C. **Na** K

D. P **Ar**

E. **Cl** Si

F. **Li** K

12. Define electronegativity. **The ability of an atom to attract electrons in a bond.**

13. What trend in electronegativity do you see as you go down a group/family on the periodic table? **Values decrease within a group. Since the valence shell becomes further away, it is harder for the nucleus to draw electrons towards it.**

14. What trend in electronegativity do you see as you go across a period/row on the periodic table? What causes this trend? **Values increase across a period because the atoms are getting smaller and the nucleus can pull electrons in easier and the atoms are closer to being like noble gases. They want to be able to complete their valence shell.**

# Periodic Trends Summary Table Answer Key

**The Princeton Review**



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