

$$\Delta G^\circ = -2.85 \text{ kJ}$$

ह) मान निकाल  $E^\circ$  की सहा

$$\Delta G^\circ = -2.85 \text{ kJ}$$

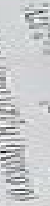
$$\log K_c = \frac{-\Delta G^\circ}{2.303RT}$$

$$= \frac{-(-2.85)}{2.303 \times 8.314 \times 298} = 0.519$$

$$\log K_c = 0.519$$

$$K_c = 3.27$$

5. निम्नलिखित सेले को प्ले X पर सेल स्क्रिप्ट पर and लिखिए



दिए गए हैं

$$E^\circ_{\text{Mg}^{2+}/\text{Mg}} = -2.37 \text{ V}$$

$$E^\circ_{\text{Cu}^{2+}/\text{Cu}} = +0.34 \text{ V}$$

$$E^\circ_{\text{Fe}^{2+}/\text{Fe}} = -0.44 \text{ V}$$

$$E^\circ_{\text{Zn}^{2+}/\text{Zn}} = -0.76 \text{ V}$$

$$E^\circ_{\text{Pb}^{2+}/\text{Pb}} = -0.13 \text{ V}$$

$$E^\circ_{\text{H}^+/\text{H}_2} = 0.00 \text{ V}$$

ह) मान निकाल  $E^\circ$  की सहा

$$\Delta G^\circ = -nFE^\circ$$

$$\log K_c = \frac{-\Delta G^\circ}{2.303RT}$$

$$= \frac{-(-2.85)}{2.303 \times 8.314 \times 298} = 0.519$$

$$\log K_c = 0.519$$

$$K_c = 3.27$$

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# Notes Chemistry For Up Board

**Rachel Sandford**



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