

Chemistry 30

Titration Analysis of Permanganate with Acidified Iron (II) Ammonium Sulfate

Name:

Group Members:

Date:

Lab Station:

Background:

In this investigation, you assume the role of a laboratory technician working in a consumer products laboratory, testing the concentration of consumer solutions. Here, you will perform a titration required to quantitatively determine the concentration of a solution of potassium permanganate by reacting it to completion with an acidified standard solution.

Purpose:

The technological purpose of this investigation is to determine the concentration of permanganate solution with a primary standard. This solution is then tested to evaluate the percent concentration of other consumer solutions.

Problem:

What is the amount concentration of potassium permanganate in solution?

Safety:

Sulfuric acid is corrosive (irritant) and can burn sulfuric acid and potassium permanganate are irritants. Avoid inhaling any solid, and avoid skin or eye contact.



The H_2SO_4 solution is concentrated strong acid; you MUST wear eye protection at all times. H_2SO_4 can cause irritation to skin and damage clothing; contact with skin or clothing should be rinsed thoroughly. If any acid is poured onto the skin or clothing, wash with water for 15 minutes. There will be no food or drinks allowed in the lab room during this lab and they must be tied back.

Spills must be reported to your teacher. All solutions can be disposed down the sink.

Design:

Your teacher has prepared a solution of potassium permanganate by dissolving 1.286 g of solid with 25 mL of distilled water. Please note that the solution you are working with in this lab has a theoretical concentration of approximately 0.010 mol/L. However, in this lab you will determine the EXACT concentration of this solution.

Samples of this diluted solution are titrated with a standardized iron(II) ammonium sulfate heptahydrate solution with a known concentration using the permanganate ions as the indicator.

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KJ Lindholm-Leary



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