



Name _____ Date _____ Block _____

Lab: Redox Half-Reactions

Purpose: To determine the oxidation and reduction half-reactions for simple single-replacement reactions.

Observations

Record your observations below (e.g., gas bubbles or turns black). If nothing occurs, write "No Rxn."

		1	2	3	4	5	6
		ZnSO ₄	H ₂ SO ₄	FeSO ₄	AgNO	CuSO ₄	MgSO ₄
A	Zn	1	2	3	4	5	6
B	Cu	7	8	9	10	11	12
C	Mg	13	14	15	16	17	18
D	Fe	19	20	21	22	23	24

Analysis

For each reaction that took place:

- Write a complete and balanced chemical equation, including phases.
- Assign oxidation numbers for every element
- Write an the oxidation and reduction half-reactions, including phases.

Example	Overall Rxn:	$\text{Ba(s)} + \text{H}_2\text{SO}_4\text{(aq)} \rightarrow \text{H}_2\text{(g)} + \text{BaSO}_4\text{(aq)}$							
	Oxidation numbers:	0	+1	+6	-2	0	+2	+6	-2
	Oxidation-half Rxn:	$\text{Ba(s)} \rightarrow \text{Ba}^{2+}\text{(aq)} + 2\text{e}^-$							
	Reduction-half Rxn:	$2\text{H}^+\text{(aq)} + 2\text{e}^- \rightarrow \text{H}_2\text{(g)}$							

Redox Lab Answer Key

Donald L. Sparks



Redox Lab Answer Key:

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