

## Polyatomic Ions WS

Ammonium ( $\text{NH}_4$ )<sup>1+</sup>  
 Carbonate ( $\text{CO}_3$ )<sup>2-</sup>  
 Chromate ( $\text{CrO}_4$ )<sup>2-</sup>

Dichromate ( $\text{Cr}_2\text{O}_7$ )<sup>2-</sup>  
 Hydroxide ( $\text{OH}$ )<sup>1-</sup>  
 Nitrate ( $\text{NO}_3$ )<sup>1-</sup>

Nitrite ( $\text{NO}_2$ )<sup>1-</sup>  
 Sulfate ( $\text{SO}_4$ )<sup>2-</sup>  
 Sulfite ( $\text{SO}_3$ )<sup>2-</sup>

Phosphate ( $\text{PO}_4$ )<sup>3-</sup>

Write the **NAME** of each of the following compounds.

1.  $(\text{NH}_4)\text{Cl}$  Ammonium Chloride
2.  $\text{Be}(\text{SO}_4)$  Beryllium Sulfate
3.  $(\text{NH}_4)_3\text{N}$  Ammonium Nitride
4.  $\text{MgCl}_2$  Magnesium Chloride
5.  $\text{NH}_4(\text{NO}_3)$  Ammonium Nitrate
6.  $\text{Sr}_3(\text{PO}_4)_2$  Strontium Phosphate
7.  $\text{Zn}(\text{CrO}_4)_2$  Zinc Dichromate
8.  $\text{K}_2(\text{Cr}_2\text{O}_7)$  Potassium Dichromate

9.  $\text{Ga}(\text{ClO}_3)_3$  Gallium Chlorate
10.  $\text{Cu}(\text{OH})$  Copper I Hydroxide
11.  $(\text{NH}_4)_3(\text{PO}_4)$  Ammonium Phosphate
12.  $\text{Fe}(\text{SO}_4)$  Iron II Sulfate
13.  $\text{Mg}(\text{NO}_2)_2$  Magnesium Nitrite
14.  $(\text{NH}_4)\text{NO}_2$  Ammonium Nitrite
15.  $\text{Na}_2(\text{Cr}_2\text{O}_7)$  Sodium Dichromate
16.  $\text{Na}(\text{OH})$  Sodium Hydroxide

Write the **CHEMICAL FORMULA** for each of the given **NAMES**.....cross charges

- |                         |                                     |   |                               |                   |
|-------------------------|-------------------------------------|---|-------------------------------|-------------------|
| 17. calcium carbonate   | $\text{Ca}^{2+} \text{CO}_3^{2-}$   | = | $\text{Ca}_2(\text{CO}_3)_2$  | = $\text{CaCO}_3$ |
| 18. barium nitrate      | $\text{Ba}^{+2} \text{NO}_3^{-1}$   | = | $\text{Ba}(\text{NO}_3)$      |                   |
| 19. ammonium sulfate    | $\text{NH}_4^{+1} \text{SO}_4^{-2}$ | = | $(\text{NH}_4)_2 \text{SO}_4$ |                   |
| 20. aluminum hydroxide  | $\text{Al}^{+3} \text{OH}^{-1}$     | = | $\text{Al}(\text{OH})_3$      |                   |
| 21. calcium phosphate   | $\text{Ca}^{+2} \text{PO}_4^{-3}$   | = | $\text{Ca}_3(\text{PO}_4)_2$  |                   |
| 22. cesium nitrate      | $\text{Cs} \text{NO}_2^{-1}$        | = | $\text{CsNO}_3$               |                   |
| 23. sodium nitrite      | $\text{Na}^{+1} \text{NO}_3^{-1}$   | = | $\text{NaNO}_3$               |                   |
| 24. calcium sulfate     | $\text{Ca}^{+2} \text{SO}_4^{-2}$   | = | $\text{CaSO}_4$               |                   |
| 25. beryllium sulfate   | $\text{Be} \text{SO}_4^{-2}$        | = | $\text{BeSO}_4$               |                   |
| 26. sodium carbonate    | $\text{Na}^{+1} \text{CO}_3^{-2}$   | = | $\text{Na}_2\text{CO}_3$      |                   |
| 27. magnesium phosphate | $\text{Mg}^{+2} \text{PO}_4^{-3}$   | = | $\text{Mg}_3(\text{PO}_4)_2$  |                   |
| 28. calcium phosphate   | $\text{Ca}^{+2} \text{PO}_4^{-3}$   | = | $\text{Ca}_3(\text{PO}_4)_2$  |                   |

# Polyatomic Ions Ws

**RS Peters**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the 'RS Peters' text.

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