

## Model 2 – Ternary Acids (Oxyacids)

| Acid                    | Name of Acid in Aqueous Solution | Cation (+)              | Polyatomic Anion (-) | Polyatomic Anion Name |
|-------------------------|----------------------------------|-------------------------|----------------------|-----------------------|
| $\text{HClO}_3$         | Chloric acid                     | $\text{H}_3\text{O}^+$  | $\text{ClO}_3^{-1}$  | Chlorate              |
| $\text{H}_2\text{SO}_3$ | Sulfurous acid                   | $2\text{H}_3\text{O}^+$ | $\text{SO}_3^{-2}$   | Sulfite               |
| $\text{H}_2\text{SO}_4$ | Sulfuric acid                    | $2\text{H}_3\text{O}^+$ | $\text{SO}_4^{-2}$   | Sulfate               |
| $\text{H}_3\text{PO}_3$ | Phosphorous acid                 | $3\text{H}_3\text{O}^+$ | $\text{PO}_3^{-3}$   | Phosphite             |
| $\text{H}_3\text{PO}_4$ | Phosphoric acid                  | $3\text{H}_3\text{O}^+$ | $\text{PO}_4^{-3}$   | Phosphate             |
| $\text{HNO}_3$          | Nitric acid                      | $\text{H}_3\text{O}^+$  | $\text{NO}_3^{-1}$   | Nitrate               |
| $\text{HNO}_2$          | Nitrous acid                     | $\text{H}_3\text{O}^+$  | $\text{NO}_2^{-1}$   | Nitrite               |
| $\text{H}_2\text{CO}_3$ | Carbonic acid                    | $2\text{H}_3\text{O}^+$ | $\text{CO}_3^{-2}$   | Carbonate             |

6. Look at the formulas of the ternary acids in Model 2.

a. How are ternary acids different from binary acids in their structure?

*Ternary acids contain three elements and they also contain the element oxygen.*

b. What number do you think the prefix "ter-" refers to?

*"ter" refers to three*

7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.

8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is "ic"

Polyatomic anion ending is "-ite" → acid name ending is "ous"

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when naming  $\text{HClO}_3$ ?

*$\text{ClO}_3^{-1}$  has the name of Chlorate. "ate" must be replaced with "ic". If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula  $\text{HCl}$ , not  $\text{HClO}_3$ .*

10. Write a rule for naming ternary acids.

*- The prefix "hydro" is not used in the name when an acid contains more than two elements and one of the elements is oxygen.*

*- "ate" is replaced with "ic" and "ite" is replaced with "ous".*

11. Predict the formula for chlorous acid.



12. Circle the acid(s) below that would be named beginning with the prefix "hydro-."



# Polyatomic Ions Pogil Worksheet Answers

**JL Elias**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

## **Polyatomic Ions Pogil Worksheet Answers:**

**Polyatomic Ions for Surface Analysis and Modification** Erick Ryan Fuoco, 2003      **The Particulate Nature of Polyatomic Ions**, 2007      Investigations Into the Origins of Polyatomic Ions in Inductively Coupled Plasma-mass Spectrometry Sally M. McIntyre, 2010      **Fragmentation of Diatomic and Polyatomic Ions in the Gas Phase** Christopher John Proctor, 1981      **Investigations Into the Origins of Polyatomic Ions in Inductively Coupled Plasma-mass Spectrometry**, 2010

An inductively coupled plasma mass spectrometer ICP MS is an elemental analytical instrument capable of determining nearly all elements in the periodic table at limits of detection in the parts per quadrillion and with a linear analytical range over 8 to 10 orders of magnitude. Three concentric quartz tubes make up the plasma torch. Argon gas is spiraled through the outer tube and generates the plasma powered by a looped load coil operating at 27.1 or 40.6 MHz. The argon flow of the middle channel is used to keep the plasma above the innermost tube through which solid or aqueous sample is carried in a third argon stream. A sample is progressively desolvated, atomized, and ionized. The torch is operated at atmospheric pressure. To reach the reduced pressures of mass spectrometers, ions are extracted through a series of two approximately one millimeter wide circular apertures set in water-cooled metal cones. The space between the cones is evacuated to approximately one torr. The space behind the second cone is pumped down to or near to the pressure needed for the mass spectrometer MS. The first cone, called the sampler, is placed directly in the plasma plume and its position is adjusted to the point where atomic ions are most abundant. The hot plasma gas expands through the sampler orifice and in this expansion is placed the second cone, called the skimmer. After the skimmer, traditional MS designs are employed, i.e., quadrupoles, magnetic sectors, time-of-flight. ICP MS is the leading trace element analysis technique. One of its weaknesses are polyatomic ions. This dissertation has added to the fundamental understanding of some of these polyatomic ions, their origins, and behavior. Although mainly continuing the work of others, certain novel approaches have been introduced here. Chapter 2 includes the first reported efforts to include high temperature corrections to the partition functions of the polyatomic ions in ICP MS. This and other objections to preceding papers in this area were addressed. Errors in the measured  $T_{\text{sub}}$  gas values were found for given errors in the experimental and spectroscopic values. The ionization energy of the neutral polyatomic ion was included in calculations to prove the validity of ignoring more complicated equilibria. Work was begun on the question of agreement between kinetics of the plasma and interface and the increase and depletion seen in certain polyatomic ions. This dissertation was also the first to report day-to-day ranges for  $T_{\text{sub}}$  gas values and to use a statistical test to compare different operating conditions. This will help guide comparisons of previous and future work. Chapter 4 was the first attempt to include the excited electronic state 2 in the partition function of ArO, as well as the first to address the different dissociation products of the ground and first electronic levels of ArO. Chapter 5 reports an interesting source of memory in ICP MS that could affect mathematical corrections for polyatomic ions. For future work on these topics, I

suggest the following experiments and investigations Clearly not an extensive list they are instead the first topics curiosity brings to mind 1 Measurement of T sub gas values when using the flow injection technique of Appendix B It was believed that there was a fundamental difference in the plasma when the auto sampler was used versus a continuous injection Is this reflected in T sub gas values 2 The work of Chapter 3 can be expanded and supplemented with more trials new cone materials i e copper stainless steel and more cone geometries Some of this equipment is already present in the laboratory others could be purchased or made 3 T sub gas values from Chapter 3 could be correlated with instrument pressures during the experiment Pressures after the skimmer cone were recorded for many days but have yet to be collated with the measured T sub gas values 4 The work in Chapter 5 could be expanded to include more metals Does the curious correlation between measured T sub gas and element boiling point persist 5 Investigate non linear correlations to T sub gas values of the MO memory in Chapter 5 Temperatures along the skimmer walls are not a linear gradient Ring deposits have been observed on the cone and photographs of the interface show light intensities shaping a sort of tailing peak along the outside skimmer wall Is there a physical property of the metals or metal oxides that would give this peak with the T sub gas values 6 Chemical state speciation of the metal deposits on the skimmers of Chapter 5 There may be a more logical correlation between Tgas and a physical property of the deposit ing chemical if all the metals do not deposit in the same form 7 A collaboration with our computational colleagues would be most welcome Newer calculations for ArO and RuO would be very helpful

**High Resolution Studies of the Origins of Polyatomic Ions in Inductively Coupled Plasma-Mass Spectrometry** Jill Wisnewski Ferguson, 2006 The inductively coupled plasma ICP is an atmospheric pressure ionization source Traditionally the plasma is sampled via a sampler cone A supersonic jet develops behind the sampler and this region is pumped down to a pressure of approximately one Torr A skimmer cone is located inside this zone of silence to transmit ions into the mass spectrometer The position of the sampler and skimmer cones relative to the initial radiation and normal analytical zones of the plasma is key to optimizing the useful analytical signal 1 The ICP both atomizes and ionizes the sample Polyatomic ions form through ion molecule interactions either in the ICP or during ion extraction 1 Common polyatomic ions that inhibit analysis include metal oxides MO sup adducts with argon the gas most commonly used to make up the plasma and hydride species While high resolution devices can separate many analytes from common interferences this is done at great cost in ion transmission efficiency a loss of 99% when using high versus low resolution on the same instrument 2 Simple quadrupole devices which make up the bulk of ICP MS instruments in existence do not present this option Therefore if the source of polyatomic interferences can be determined and then manipulated this could potentially improve the figures of merit on all ICP MS devices not just the high resolution devices often utilized to study polyatomic interferences

**Standard Thermodynamic Functions of Gaseous Polyatomic Ions at 100-1000 K** Aharon Loewenschuss, Y. Marcus, 1987

*Fragmentation of Diatomic and Polyatomic Ions in the Gas Phase* Christopher John Proctor, 1981 *Unimolecular and*

*Collision-induced Dissociation Study of Polyatomic Ions at High Collision Energy* Xuedong Zhou, 2001      **Polyatomic Ion Impact on Solids and Related Phenomena** Y Le Beyec, Y Hoppilliard, H Bernas, 1994      *Polyatomic Ion-surface Interactions* Luke Hanley, 1998      **Experimental Studies of Polyatomic Ion Interactions with Clean and Adsorbate Covered Metal Surfaces** Samuel B. Wainhaus, 1997      **Special issue polyatomic ion surface interactions** Luke Hanley, 1998      **Determination of Optimum Conditions for Distinguishing the Pulse Height Distributions of Atomic and Polyatomic Ions** M. J. Kristo, 2006

This work explored the use of pulse height distributions PHD from multiplier type detectors as a means of detecting and eliminating the effects of polyatomic interferences in secondary ion mass spectrometry SIMS We explored the behavior of PHD for sup 235 U sup sup 208 Pb sup 27 Al sup and sup 207 Pb sup 28 Si sup all with a nominal mass to charge ratio of 235 In every case the distribution for the atomic ion sup 235 U sup was clearly shifted relative to the distributions for sup 208 Pb sup 27 Al sup and sup 207 Pb sup 28 Si sup When the first surface of the detector is metallic in character the polyatomic ions are shifted to larger pulse heights relative to the atomic ion When the first surface of the detector is oxide in character the atomic ion is shifted to larger pulse heights relative to the polyatomic ions The relative positioning appear to be stable for a given detector over time at the same secondary ion impact energy Consequently it appears to be feasible to use PHD data to detect interfering polyatomic ions and eliminate their deleterious effects using peak deconvolution techniques Consequently the updated Ultrafast RAE detector will be designed to make the pulse height information available to the data acquisition system      **High Resolution Studies of the Origins of Polyatomic Ions in Inductively Coupled Plasma-Mass Spectrometry** , 2006

The inductively coupled plasma ICP is an atmospheric pressure ionization source Traditionally the plasma is sampled via a sampler cone A supersonic jet develops behind the sampler and this region is pumped down to a pressure of approximately one Torr A skimmer cone is located inside this zone of silence to transmit ions into the mass spectrometer The position of the sampler and skimmer cones relative to the initial radiation and normal analytical zones of the plasma is key to optimizing the useful analytical signal 1 The ICP both atomizes and ionizes the sample Polyatomic ions form through ion molecule interactions either in the ICP or during ion extraction 1 Common polyatomic ions that inhibit analysis include metal oxides MO adducts with argon the gas most commonly used to make up the plasma and hydride species While high resolution devices can separate many analytes from common interferences this is done at great cost in ion transmission efficiency a loss of 99% when using high versus low resolution on the same instrument 2 Simple quadrupole devices which make up the bulk of ICP MS instruments in existence do not present this option Therefore if the source of polyatomic interferences can be determined and then manipulated this could potentially improve the figures of merit on all ICP MS devices not just the high resolution devices often utilized to study polyatomic interferences      **Calculated Angular Distribution of Fragment Ions from Polyatomic-ion Collisions** , 1982

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**Negative Ions** William B. Clodius, 1982

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## **Table of Contents Polyatomic Ions Pogil Worksheet Answers**

1. Understanding the eBook Polyatomic Ions Pogil Worksheet Answers
  - The Rise of Digital Reading Polyatomic Ions Pogil Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Polyatomic Ions Pogil Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Polyatomic Ions Pogil Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Polyatomic Ions Pogil Worksheet Answers
  - Personalized Recommendations
  - Polyatomic Ions Pogil Worksheet Answers User Reviews and Ratings
  - Polyatomic Ions Pogil Worksheet Answers and Bestseller Lists
5. Accessing Polyatomic Ions Pogil Worksheet Answers Free and Paid eBooks

- Polyatomic Ions Pogil Worksheet Answers Public Domain eBooks
  - Polyatomic Ions Pogil Worksheet Answers eBook Subscription Services
  - Polyatomic Ions Pogil Worksheet Answers Budget-Friendly Options
6. Navigating Polyatomic Ions Pogil Worksheet Answers eBook Formats
- ePub, PDF, MOBI, and More
  - Polyatomic Ions Pogil Worksheet Answers Compatibility with Devices
  - Polyatomic Ions Pogil Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Polyatomic Ions Pogil Worksheet Answers
  - Highlighting and Note-Taking Polyatomic Ions Pogil Worksheet Answers
  - Interactive Elements Polyatomic Ions Pogil Worksheet Answers
8. Staying Engaged with Polyatomic Ions Pogil Worksheet Answers
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Polyatomic Ions Pogil Worksheet Answers
9. Balancing eBooks and Physical Books Polyatomic Ions Pogil Worksheet Answers
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Polyatomic Ions Pogil Worksheet Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Polyatomic Ions Pogil Worksheet Answers
- Setting Reading Goals Polyatomic Ions Pogil Worksheet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Polyatomic Ions Pogil Worksheet Answers
- Fact-Checking eBook Content of Polyatomic Ions Pogil Worksheet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development



- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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