

NJXDG EXAM 2023 STUDY GUIDE WITH COMPLETE SOLUTION

Applying topical fluoride gel or foam (efda) - Selecting the tray

1) Select a disposable tray that is of the appropriate size for the patient's mouth. The tray must be long and sufficiently deep to cover all erupted teeth completely without extending beyond the distal surface of the most posterior tooth. Trays are available in sizes to fit primary, mixed, and adult dentition. If the patient's mouth can accommodate it, you may use a double-arch tray. This saves time by treating both arches at the same time. Remember, trays are discarded after a single use & if you try a tray in the mouth but do not use it that tray must be discarded.

Preparing the teeth

2) Check to see whether calculus is present; if it is not, no preparation is required. Fluoride diffuses easily through the acquired pellicle and bacterial plaque.
3) If calculus is present, request that the dentist or dental hygienist remove it. The presence of plaque will not affect the uptake of fluoride.

Applying the topical fluoride

4) Seat the patient in an upright position and explain the procedure. Having the pt upright prevents gel from going into the throat.
5) Instruct the pt not to swallow the fluoride.
6) Select the appropriate tray, and load it with a minimal amount of fluoride, following guidelines according to the patient's age.
7) Dry the teeth using air from the air water syringe. For fluoride to be maximally effective, the teeth must be dry when the fluoride is applied.
8) Insert the tray, and place cotton rolls between the arches. Ask the patient to bite up and down gently on the cotton rolls, to squeeze the fluoride over all tooth surfaces.
9) promptly place the saliva ejector, and tilt the patient's head forward.
10) Set the timer for the appropriate amount of time in accordance with the manufacturer's instructions. During this time, do not leave the patient unattended.
11) On completion, remove the tray, but do not allow the pt to rinse or swallow. Promptly use the saliva ejector or the high-volume oral evaluator tip to close the lips tightly around the saliva ejector. Removing excess saliva and fluoride solution will make the pt more comfortable and less likely to rinse with water.
12) Instruct and the pt not to rinse, eat, drink, or brush the teeth for at least 30 minutes. These activities could disturb the action of the fluoride.

Documentation

- date
- "applied APF fluoride gel. Instructed pt not to eat for 30 min.)
- operator

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Liyong Dong



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the boiling point of an organic liquid is the temperature at which its vapour pressure equals the atmospheric pressure over the liquid or it is the temperature at which the vapour and liquid phases are in equilibrium at agiven pressure the boiling point is considered as a criterion of purity of a compound and is useful for identification of

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given that the boiling points of ethanol and methanol 78 c and 64 c respectively are far off in a significant way this seems like a valid conclusion conclusion thermometer calibrations are necessary in order to collect reliable information as precise calculations are imperative to gathering correct results

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melting point 1 a push the open end of a capillary tube into the powdered acetamide b move the powder to the closed end of the capillary tube by tapping it on the table repeat until the the powdered acetamide occupies 1 2 mm of the capillary tube end c

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discussion and conclusion the technique recrystallization was used because it is a simple method to purify a solid compound the percentage yield for benzoic acid and acetanilide acid being high indicated the initial sample had a low concentration of impurities with the melting point being in a specific

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determination of boiling points introduction the boiling point of a compound is the temperature at which it changes from a liquid to a gas this is a physical property often used to identify substances or to check the purity of the compound it

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conclusion after completing this experiment the melting point of three different unknown compounds was obtained using the mel temp apparatus compound a had a melting point range of 114 6 115 5 c thus revealing it to be acetanilide compound b had a melting point range of 122 6 124 5 c thus

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melting point can be used as the identity and purity of an organic solid however it is not the same as the boiling point the boiling point is a characteristic unique of pure liquids the boiling point is the temperature at which the vapor pressure of a liquid is equal to the atmospheric pressure

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aug 11 2020 boiling points are very sensitive to changes in applied pressure so all boiling points should be reported with the measured pressure a compound s normal boiling point refers to its boiling point at a pressure of 760 text mm ce hg

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the purpose of the experiment is to understand what vapour pressure is and how it changes with temperature to understand the relationship between vapour pressure at a given temperature and boiling temperature at a given pressure to determine the vapour pressure of a pure liquid at various temperatures

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exploring boiling points abstract in this experiment we are given a set of straight chain alkanes and a set of compounds to compare by following the trendline of the straight chain alkanes it could be concluded that as the molar mass of the straight chain alkanes increases their boiling points increase as well

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laboratory partners samantha and mackenzie goals □ correctly determine the boiling point of an unknown liquid □ measure the refractive index of an unknown □ use the gas chromatogram to determine the mass percentage of a mixture and

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50 c so in order to convert an experimental measurement taken in calgary which is at higher altitude than sea level and so is at a lower pressure to that reported for sea level higher pressure one needs to add a correction factor since the boiling point at sea level is higher than that at higher altitudes

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introduction in chemistry one may encounter an unknown substance that need to be determined by physical properties such as boiling point melting point density solubility etc in this experiment i am to figure out the boiling point of a mystery substance in order to identify it chemical responsibility this experiment requires the handling of a bunsen burner boiling
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there are a variety of methods by which a sample s boiling point can be determined including distillation reflux and by using a thiele tube the most straightforward method uses a thiele tube and has the advantage of using less than 0 5ml 0 5 ml of material

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