

CHEMISTRY



Properties of the Periodic Table

- Elements are arranged in order of increasing atomic number.
- Elements in the same group have similar chemical properties.
- The periodic table is divided into four main blocks: s-block, p-block, d-block, and f-block.
- The periodic table is a useful tool for predicting the properties of elements.

Atomic Trends

- Atomic size increases from top to bottom and from right to left.
- Ionization energy increases from bottom to top and from left to right.
- Electronegativity increases from bottom to top and from left to right.

Atomic Structure

Atoms are made up of three main parts: protons, neutrons, and electrons. Protons and neutrons are found in the nucleus, while electrons orbit the nucleus in shells.

Atomic Number

The atomic number of an element is the number of protons in its nucleus. It is represented by the symbol Z .

Mass Number

The mass number of an element is the total number of protons and neutrons in its nucleus. It is represented by the symbol A .

Atomic Structure Diagram

Atomic Structure Table

Element	Atomic Number (Z)	Mass Number (A)	Protons	Neutrons	Electrons
Hydrogen	1	1	1	0	1
Helium	2	4	2	2	2
Lithium	3	7	3	4	3
Boron	5	11	5	6	5
Carbon	6	12	6	6	6
Nitrogen	7	14	7	7	7
Oxygen	8	16	8	8	8
Fluorine	9	19	9	10	9
Neon	10	20	10	10	10

Atomic Structure Diagram

Atomic Structure Table

Element	Atomic Number (Z)	Mass Number (A)	Protons	Neutrons	Electrons
Sodium	11	23	11	12	11
Magnesium	12	24	12	12	12
Aluminum	13	27	13	14	13
Silicon	14	28	14	14	14
Phosphorus	15	31	15	16	15
Sulfur	16	32	16	16	16
Chlorine	17	35	17	18	17
Argon	18	36	18	18	18

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