

Honors Chemistry Quiz Chapter 6: Thermochemistry

Name: _____ Period: _____ Date: _____

This quiz is worth 40 points; each correct response is 2 points. Only those quizzes completed in black ink will be graded. Good luck!

1) Energy that is associated with the position or composition of an object is called

- A) kinetic energy
- B) thermal energy
- C) potential energy
- D) chemical energy
- E) None of the above.

2) Which of the following signs on q and w represent a system that is doing work on the surroundings, as well as losing heat to the surroundings?

- A) $q = -$, $w = -$
- B) $q = +$, $w = +$
- C) $q = -$, $w = +$
- D) $q = +$, $w = -$
- E) None of these represent the system referenced above.

3) Calculate the change internal energy (ΔE) for a system that is giving off 45.0 kJ of heat and is performing 855 J of work on the surroundings.

- A) 44.1 kJ
- B) -44.1 kJ
- C) -45.9 kJ
- D) 9.00×10^2 kJ
- E) -9.00×10^2 kJ

4) Define specific heat capacity.

- A) the quantity of heat required to raise the temperature of 1 mole of a substance by 1°C
- B) the quantity of heat required to change a system's temperature by 1°C
- C) the quantity of heat required to raise the temperature of 1 gram of a substance by 1°C
- D) the quantity of heat required to raise the temperature of 1 gram of a substance by 1°F
- E) the quantity of heat required to raise the temperature of 1 liter of a substance by 1°C

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