


- (c) 350 cm, convergent
(d) 150 cm, divergent
16. Three resistances A, B and C each of 5Ω are connected to a battery of 10 V as shown in the figure. Calculate the current through C:
- 
- (a) 2 amp (b) 1 amp
(c) 3 amp (d) 0.6 amp
17. The device used to generate electrical energy is:
- (a) Electric motor (b) Generator
(c) Galvanometer (d) Voltmeter
18. The number of molecules in 11 g of CO_2 are:
- (a) 0.25×10^{23} (b) 0.50×10^{23}
(c) 1.00×10^{23} (d) 1.51×10^{23}
19. The value of charge / mass ratio of electron was determined by:
- (a) W.K. Roentgen (b) J.J. Thomson
(c) Marie Curie (d) Niels Bohr
20. Which of the following reactions is an example of combination reaction?
- (a) $\text{Zn(s)} + \text{CuSO}_4(\text{aq}) \rightarrow \text{ZnSO}_4(\text{aq}) + \text{Cu(s)}$
(b) $\text{BaCl}_2(\text{aq}) + \text{Na}_2\text{SO}_4(\text{aq}) \rightarrow \text{BaSO}_4(\text{s}) + 2\text{NaCl}(\text{aq})$
(c) $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO(s)} + \text{CO}_2(\text{g})$
(d) $\text{C(s)} + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g})$
21. The element 'X' and 'Y' have atomic numbers 12 and 17 respectively. Element 'X' reacts with element 'Y' to form compound with molecular formula:
- (a) XY (b) XY_2
(c) XY_3 (d) X_3Y_3
22. The atomic radius (pm) of Li, Na, K and Rb varies in the order:
- (a) $\text{Na} < \text{K} < \text{Rb} < \text{Li}$ (b) $\text{K} < \text{Na} < \text{Li} < \text{Rb}$
(c) $\text{Li} < \text{Na} < \text{K} < \text{Rb}$ (d) $\text{Rb} < \text{K} < \text{Na} < \text{Li}$
23. The electron affinity value (kJ mol^{-1}) of fluorine (F) is less than:
- (a) Hydrogen (H) (b) Lithium (Li)
(c) Oxygen (O) (d) Chlorine (Cl)
24. Which of the following is an example of strong electrolyte?
- (a) H_2CO_3 (b) NH_4OH
(c) NaCl (d) $\text{COOH} - \text{COOH}$
25. If pH of a solution changes from 5 to 4, the change in hydrogen ion concentration shall be:
- (a) two times (b) five times
(c) ten times (d) twenty times
26. Bleaching powder is manufactured by the reaction of:
- (a) CaCl_2 and CaCO_3 (b) Cl_2 and Ca(OH)_2
27. (c) Cl_2 and CaSO_4 (d) Cl_2 and $\text{Ca(HCO}_3)_2$
The hard glass is obtained by fusing:
- (a) soda ash, sand and limestone
(b) a mixture of sand, lime, borax and alkali carbonates
(c) potassium carbonate and limestone
(d) potassium carbonate, lead oxide and sand
28. Which of the following metals is most reactive?
- (a) Aluminium (b) lead
(c) Mercury (d) Silver
29. Which of the following alloys contains chromium?
- (a) Steel (b) Stainless steel
(c) Magnadium (d) Brass
30. Which of the following is monomer of natural rubber?
- (a) Chloroethene (b) Chloroprene
(c) Isoprene (d) Buta-1, 3-diene
31. Alkaline KMnO_4 oxidises propanone to:
- (a) propanoic acid (b) ethanoic acid
(c) methanoic acid (d) oxalic acid
32. Which of the following metals can displace Zn from ZnSO_4 solution?
- (a) Calcium (b) Copper
(c) Iron (d) Mercury
33. Heating of sodium ethanoate with soda lime yields:
- (a) ethane (b) methane
(c) ethanol (d) methanol
34. The compound formed by the reaction of ethyne with bromine is:
- (a) $\text{Br}-\text{CH}=\text{CH}-\text{Br}$ (b) $\text{Br}-\text{CH}_2-\text{CHBr}_2$
(c) $\text{CH}_2=\text{CH}-\text{Br}$ (d) $\text{Br}_2\text{CH}-\text{CHBr}_2$
35. Cell were first discovered by Robert Hooke in the year:
- (a) 1665 (b) 1674 (c) 1831 (d) 1839
36. The plant cells have a rigid cell wall that lies:
- (a) Outside the plasma membrane
(b) Inside the plasma membrane
(c) In between the plasma membranes
(d) None of the above
37. During mitosis nucleolus and nuclear membrane are lost in stage:
- (a) Prophase (b) Metaphase
(c) Anaphase (d) Telophase
38. An example of simple and permanent tissue is:
- (a) Xylem (b) Phloem
(c) Sclerenchyma (d) All the above
39. The shape of squamous epithelial tissue is:
- (a) Cubical (b) Flattened
(c) Pillar like (d) None of the above
40. Fresh water sponge belongs to phylum:
- (a) Arthropoda (b) Annelida

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M Mark



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