

**ICSE Board
Class X Chemistry
Board Paper – 2012**

Time: 1½ hrs

Total Marks: 80

General Instructions:

1. Answers to this paper must be written on the paper provided separately.
2. You will NOT be allowed to write during the first 15 minutes. This time is to be spent in reading the question paper.
The time given at the head of this paper is the time allowed for writing the answers.
This question paper is divided into two sections.
3. **Section I** contains one question with parts (a) to (h); all the eight parts are to be answered.
4. **Section II** contains six questions numbered 2 to 7. You are to answer any four of these questions.
The intended marks of questions or for parts of questions are given in brackets [].

SECTION I (40 Marks)

Attempt all questions from this section.

Question 1

(a) Name the gas in each of the following:

[5]

- i. The gas evolved on reaction of aluminium with boiling concentrated caustic alkali solution.
- ii. The gas produced when excess ammonia reacts with chlorine.
- iii. A gas which turns acidified potassium dichromate clear green.
- iv. The gas produced when copper reacts with concentrated nitric acid.
- v. The gas produced on reaction of dilute sulphuric acid with a metallic sulphide.

(b) State one observation for each of the following:

[5]

- i. Excess ammonium hydroxide solution is added to lead nitrate solution.
- ii. Bromine vapours are passed into a solution of ethyne in carbon tetrachloride.
- iii. A zinc granule is added to copper sulphate solution.
- iv. Zinc nitrate crystals are strongly heated.
- v. Sodium hydroxide solution is added to ferric chloride solution at first a little and then in excess.

Question Paper 2011 Icse Board Of Chemistry

Michal Rosen-Zvi



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eg 1 $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$ for which $\Delta H = -57\text{kJ mol}^{-1}$ eg 2 $\text{HNO}_3\text{(aq)} + \text{KOH(aq)} \rightarrow \text{KNO}_3\text{(aq)} + \text{H}_2\text{O(l)}$ for which $\Delta H = -57\text{kJ mol}^{-1}$

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web the enthalpy change of neutralisation is defined as the enthalpy change when solutions of an acid and alkali react together to produce one mole of water for example $\text{HCl(aq)} + \text{NaOH(aq)} \rightarrow \text{NaCl(aq)} + \text{H}_2\text{O(aq)}$ you can see that one mole of water is produced when one mole of hcl or one mole of naoh reacts