

## Chemistry – Gas Laws ACT Questions [USE TO STUDY FOR YOUR TEST]

### Multiple Choice

Identify the choice that best completes the statement or answers the question.

1. At  $22^{\circ}\text{C}$ , the air pressure in a car tire is 1293 torr. Convert this pressure to atmospheres (atm).
- 12.76
  - 5.262
  - 1.701
  - 1.293

2. This diagram shows a closed-end mercury (Hg) barometer.



When atmospheric pressure is 1.30 atm, what is the height ( $h$ ), in millimeters, of the Hg in the barometer?

- 132
  - 585
  - 760
  - 988
3. A chemist compresses a 1.5 L sample of carbon dioxide ( $\text{CO}_2$ ) gas at 1.2 atm to a final volume of 0.75 L. Assuming the temperature is constant, what is the final pressure, in atm, exerted by the  $\text{CO}_2$  gas?
- 0.60
  - 0.94
  - 1.7
  - 2.4
4. A vessel connected to an open-end mercury (Hg) manometer contains nitrogen ( $\text{N}_2$ ) gas. The atmospheric pressure is 0.993 atm. The height difference between the two arms of the manometer



is 122 mm.

What is the pressure, in mm Hg, of the  $\text{N}_2$  gas?

- 355
- 633

- 755
- 877

5. At 1.00 atm, a sealed weather balloon contains 20.0 L of helium (He) gas at  $25.0^{\circ}\text{C}$ . Assume that none of the He escapes and the pressure is constant. What is the volume, in liters, of He in the weather balloon at  $35.0^{\circ}\text{C}$ ?
- 40.0
  - 28.0
  - 20.7
  - 14.3
6. At 1254 m above sea level, the city of Helena, Montana, has an average atmospheric pressure of 0.859 atm. A chemistry student in Helena finds that pure distilled water boils at  $95.9^{\circ}\text{C}$ . Which statement explains the student's results at this high elevation?
- The boiling point of water increases due to lower atmospheric pressure.
  - The boiling point of water increases due to higher atmospheric pressure.
  - The boiling point of water decreases due to lower atmospheric pressure.
  - The boiling point of water decreases due to higher atmospheric pressure.
7. For a chemistry homework assignment, Beth must consider the effusion rates of methane ( $\text{CH}_4$ ) gas and nitrogen trifluoride ( $\text{NF}_3$ ) gas at  $25^{\circ}\text{C}$ . Effusion occurs as gas particles escape from a container through a small hole. Which statement accurately compares the effusion rates of  $\text{CH}_4$  and  $\text{NF}_3$  at  $25^{\circ}\text{C}$ ?
- $\text{CH}_4$  effuses 2 times faster than  $\text{NF}_3$ .
  - $\text{CH}_4$  effuses 4 times faster than  $\text{NF}_3$ .
  - $\text{NF}_3$  effuses 2 times faster than  $\text{CH}_4$ .
  - $\text{NF}_3$  effuses 4 times faster than  $\text{CH}_4$ .
8. At standard temperature and pressure, a chemist burns 3.50 L of pentane ( $\text{C}_5\text{H}_{12}$ ) in excess oxygen ( $\text{O}_2$ ) gas.
- $$\text{C}_5\text{H}_{12}(l) + 8 \text{O}_2(g) \rightarrow 5 \text{CO}_2(g) + 6 \text{H}_2\text{O}(l)$$
- What is the maximum number of liters of carbon dioxide ( $\text{CO}_2$ ) that the reaction can produce?
- 17.5
  - 8.50
  - 5.60
  - 0.700

# Sample Test Questions On Gas Laws

**Guillaume Favre**



## **Sample Test Questions On Gas Laws:**

## Reviewing **Sample Test Questions On Gas Laws**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Sample Test Questions On Gas Laws**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve in to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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