Unit 1 Organic Chemistry

ARE YOU READY?

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Understanding Concepts

- 1. (a) 3.3-dimethy/pentane
- (b) 1-methyl-4-peopyleyelohexane
- (a) CH₃CH₂CH₂CH₃CH₃
 - (b) CH₃

CH₂CCH₂CH₂CH₂CH₂CH₃CH₃

CH₃



CH₂CH₃

CH2CHCH2CH2CH=CH2

(e) CH,CH,C=CH

3. C₂H₁₆ + 11 O₂ → 7 CO₂ + 8 H₂O

 Cyclohexane (a) and 2,3-dimethyl-2-butene (d) are structural isomers; cyclohexene (b) and hexyne (c) are structural isomers.

5. Pentane would have a higher boiling point because it has a long carbon chain allowing more van der Waals attractions between molecules. The 2,2-dimethylpropane has a more spherical shape, with fewer intermolecular attractions. More energy is required to separate molecules with stronger intermolecular attractions, and thus, these compounds have higher boiling points.

2,2-dihydroxyethane, (a), will have a greater solubility in water than does methylbenzene, (b).

7. H, (C, S), (N, CI), O

8. (a) intramolecular: single covalent bonds; intermolecular: van der Waals

(b) intramolecular: single covalent bonds; intermolecular: van der Waals forces and hydrogen bonds

(c) intramolecular: single covalent bonds; intermolecular: van der Waals forces and hydrogen bonds

Liquid 1 is 2-methylbutane; liquid 2 is 2-methyl-2-butene; liquid 3 is pentane.

10. (a) The safety precautions needed are as follows: wear eye protection and a lab apron; avoid direct contact with skin (if in contact with skin, wash with copious amounts of water); when diluting concentrated acids, add acid to water; handle concentrated acids in a fume hood; store in glass or plastic container in secure location.

(b) The safety precautions needed are as follows: wear eye protection and a lab apron; do not use with open flames; store in secure location.

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Camilla Rothe

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