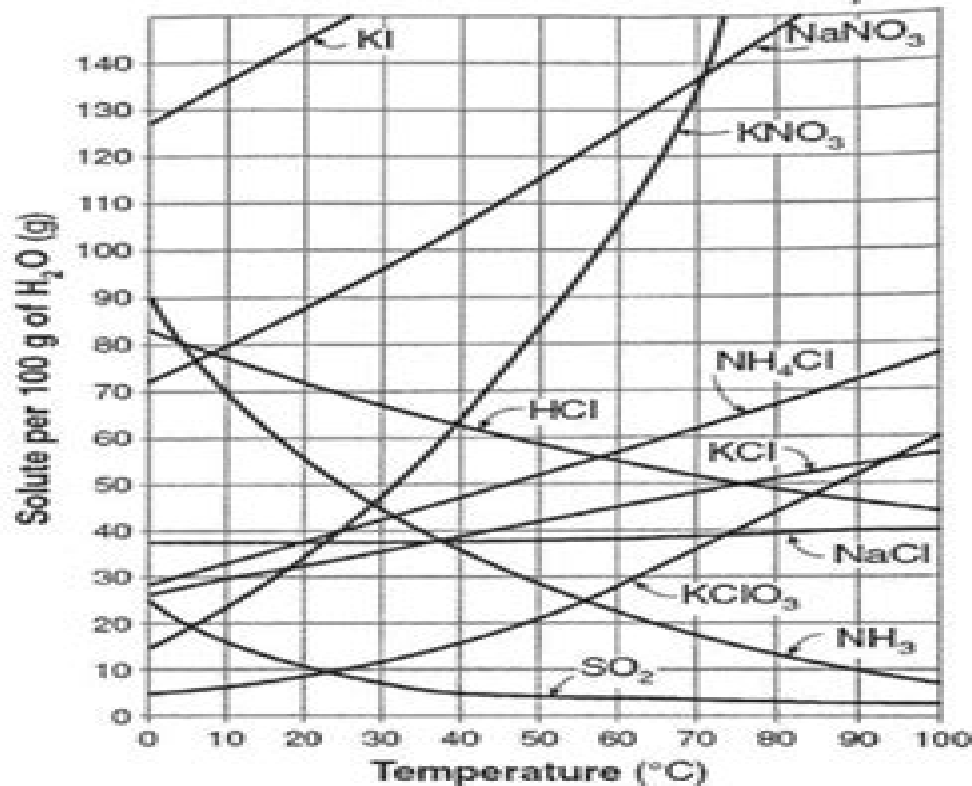


## Reading a Solubility Chart

- 1) The curve shows the # of grams of solute in a saturated solution containing 100 mL or 100 g of water at a certain temperature.
- 2) Any amount of solute below the line indicates the solution is unsaturated at a certain temperature
- 3) Any amount of solute above the line in which all of the solute has dissolved shows the solution is supersaturated.
- 4) If the amount of solute is above the line but has not all dissolved, the solution is saturated and the # grams of solute settled on the bottom of the container = total # g in solution - # g of a saturated solution at that temperature. (according to the curve)
- 5) Solutes whose curves move upward w/ increased temperature are typically solids b/c the solubility of solids increases w/ increased temperature.
- 6) Solutes whose curves move downward w/ increased temperature are typically gases b/c the solubility of gases decreases with increased temperature.



## Solubility Problems

1. Which of the salts shown on the graph is the most soluble in water at 10°C? **KI**
2. At what temperature is 90g of NaNO<sub>3</sub> a saturated solution in 100grams of water? **≈ 23°C**
3. A solution has 70grams of KNO<sub>3</sub> and is unsaturated, what is the minimum temperature? **46°C**
4. At 20°C, a saturated solution of KClO<sub>3</sub> is created in water. How many grams of potassium chlorate must be added to saturate the solution at 50°C? **12g**
5. Which of the salts shown on the graph is the least soluble in water at 10°C? **KClO<sub>3</sub>**
6. Which of the salts shown on the graph has the greatest increase in solubility as the temperature increases from 30 degrees to 60 degrees? **KNO<sub>3</sub>**

# Reading The Solubility Chart Problems Answers With Work

**Julia Schneider**



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