

Report Table (B) B: Part C Data: pH Measurements of Dilute Solutions of Acetic Acid

Solution	pH using pH paper
0.10 M $\text{HC}_2\text{H}_3\text{O}_2$	9.0 0.1 0.2 0.3 0.4 0.5 0.6 0.7 0.8 0.9
0.010 M $\text{HC}_2\text{H}_3\text{O}_2$	9.0 0.1 0.2 0.3 0.4 0.5 0.6 0.7 0.8 0.9
0.0010 M $\text{HC}_2\text{H}_3\text{O}_2$	9.0 0.1 0.2 0.3 0.4 0.5 0.6 0.7 0.8 0.9

Write the net ionic equation describing the dissociation of acetic acid in aqueous solution.

Calculate $[\text{H}^+]$ and $[\text{C}_2\text{H}_3\text{O}_2^-]$ for each concentration above, showing a sample calculation here.

$$\begin{aligned}
 \text{pH} &= -\log [\text{H}^+] & [\text{H}_2\text{PO}_4^-] &= 0.10 \text{ M} \\
 9.0 &= -\log x & x &= [0.10 - 0.0001] \text{ M} = 0.0999 \text{ M} \\
 -0.9 &= -\log x & K &= \frac{[0.0001 \text{ M}] [0.0001 \text{ M}]}{[0.0999 \text{ M}]} \\
 x &= 10^{-0.9} & &= \frac{0.0001 \text{ M}}{0.0999 \text{ M}} \\
 x &= 0.0001 \text{ M} & &= 1.0 \times 10^{-4} \text{ M}
 \end{aligned}$$

Write the equilibrium constant expression for acetic acid and calculate a K value and an apparent percent dissociation for each concentration, showing sample calculations here.

$$\begin{aligned}
 K_a &= \frac{[\text{H}^+][\text{C}_2\text{H}_3\text{O}_2^-]}{[\text{HC}_2\text{H}_3\text{O}_2]} = \frac{(1.0 \times 10^{-4})^2}{0.0999} = 1.0 \times 10^{-8} \\
 \frac{[0.0001 + 0.0001]}{0.0999} &= 0.0002 = 0.2\%
 \end{aligned}$$

What is the effect of dilution on the percent dissociation? Why should this behavior be expected?

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