



Academic Year 2019- 2020
First Term
Chemistry revision sheet

Name: _____

Date: _____

Grade: 11

Section: _____

Q1. Circle the correct options

- What is the specific heat capacity of a substance if 2.41×10^4 J are needed to change the temperature of 105.0 g of it from 25.0°C to 250.0°C ?
A. 1.02×10^{-4} J/g $^\circ\text{C}$ C. 9.18×10^{-4} J/g $^\circ\text{C}$
B. 0.918 J/g $^\circ\text{C}$ D. 1.02 J/g $^\circ\text{C}$
- Which statement about enthalpy is true?
A. Heat is given off to the surroundings in endothermic reactions.
B. Some substances have a negative specific heat capacity.
C. Specific heat capacity is the same for all liquids.
D. The sign of ΔH is always negative in exothermic reactions.
- What happens to the value of ΔH for a thermochemical reaction if the reaction is reversed?
A. ΔH has the same numerical value, and the sign changes.
B. ΔH has the same numerical value, and the sign remains the same.
C. ΔH is the reciprocal of the original value, and the sign changes.
D. ΔH is the reciprocal of the original value, and the sign remains the same.
- Which statement correctly describes an endothermic chemical reaction?
A. The products have higher potential energy than the reactants, and the ΔH is negative.
B. The products have higher potential energy than the reactants, and the ΔH is positive.
C. The products have lower potential energy than the reactants, and the ΔH is negative.
D. The products have lower potential energy than the reactants, and the ΔH is positive.

5. Given the thermochemical equation:



$$\Delta H = 114 \text{ kJ}$$

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JE Gale



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