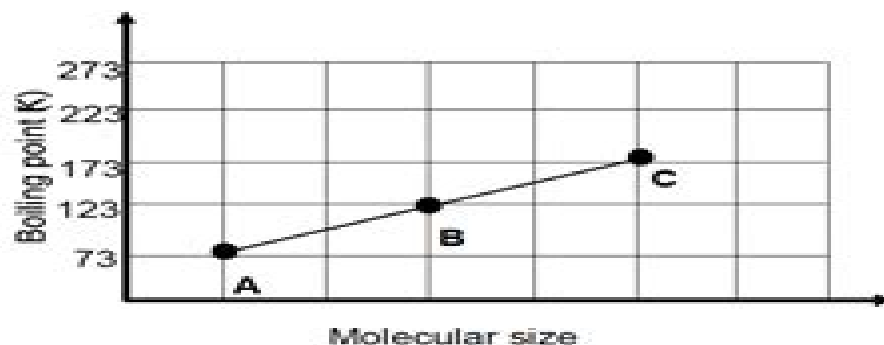


QUESTION 2

- 2.1 The graph of molecular size versus the boiling point is given below.
The letters **A**, **B** and **C** represent the compounds CH_4 , C_2H_6 and C_3H_8 respectively.



- 2.1.1 Define the term *boiling point*. (2)
- 2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph. (2)
- 2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)
- 2.1.4 Which ONE of the compounds (**A**, **B** or **C**) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph. (2)
- 2.2 Consider the two molecules in the table below.

NAME OF SUBSTANCE	FORMULA	MOLACULAR MASS (g)	BOILING POINT ($^{\circ}\text{C}$)
Ammonia	NH_3	17	- 33
Phosphine	PH_3	x	- 87,4

- 2.2.1 Determine the molecular mass (x) of phosphine (1)
- 2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

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1. Understanding the eBook Physical Science Grade 1term Test Question Paper 2caps
 - The Rise of Digital Reading Physical Science Grade 1term Test Question Paper 2caps
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Grade 1term Test Question Paper 2caps
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Grade 1term Test Question Paper 2caps
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Grade 1term Test Question Paper 2caps
 - Personalized Recommendations
 - Physical Science Grade 1term Test Question Paper 2caps User Reviews and Ratings
 - Physical Science Grade 1term Test Question Paper 2caps and Bestseller Lists
5. Accessing Physical Science Grade 1term Test Question Paper 2caps Free and Paid eBooks
 - Physical Science Grade 1term Test Question Paper 2caps Public Domain eBooks
 - Physical Science Grade 1term Test Question Paper 2caps eBook Subscription Services
 - Physical Science Grade 1term Test Question Paper 2caps Budget-Friendly Options
6. Navigating Physical Science Grade 1term Test Question Paper 2caps eBook Formats

- ePub, PDF, MOBI, and More
 - Physical Science Grade 1term Test Question Paper 2caps Compatibility with Devices
 - Physical Science Grade 1term Test Question Paper 2caps Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Physical Science Grade 1term Test Question Paper 2caps
 - Highlighting and Note-Taking Physical Science Grade 1term Test Question Paper 2caps
 - Interactive Elements Physical Science Grade 1term Test Question Paper 2caps
8. Staying Engaged with Physical Science Grade 1term Test Question Paper 2caps
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Grade 1term Test Question Paper 2caps
9. Balancing eBooks and Physical Books Physical Science Grade 1term Test Question Paper 2caps
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Grade 1term Test Question Paper 2caps
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Grade 1term Test Question Paper 2caps
- Setting Reading Goals Physical Science Grade 1term Test Question Paper 2caps
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Grade 1term Test Question Paper 2caps
- Fact-Checking eBook Content of Physical Science Grade 1term Test Question Paper 2caps
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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