		gerangskik sis om die hoeke tussen hulle te maksimeer. Strukturele elektronpare is bindingspare plus alleenpare. OF Valensdop-elektronpaarafstotingsmodel is 'n model om die vorms van molekules te voorspels waarin die elektronpaar vanaf die buitenste dops van 'n verwysingsatoom rondom hierdie atoom gerangskik is om die afstoting tussen hulle te minimaliseer.	
	1.2.4	The molecule is a tetrahedron [™] Die molekule is 'n tetraëder [™] 1 mark/punt	LO2
	1.2.5	Name trigonal pyramid Noem trigonale piramidale 3 marks/punte	LO4
1.3	1.3.1	The electron pair is equally shared between the two CI atoms and the resulting Cl ₂ molecule has no degree of polarity. Die elektronpaar word ewe veel gedeel tussen die twee CI-atome en die resulterende CI2-molekule het geen mate van polariteit nie. 2 marks/punte	LO2
	1.3.2	The tendency of an atom to pull the bonded pair of electrons towards itself. of Die neiging of van 'n atoom om die gebinde elektronpare na homself te trek of a marks/punte	LO2
	1.3.3	Electronegativity of Na = 6 Electronegativity of Cl = 3 Difference is 3 With a difference of greater than 2.1 electron transfer will take place and the bond elis ionic. (value of 2.1 does not need to be given for this mark) Elektronegatiwiteit van Na = 6 Elektronegatiwiteit van Cl = 3 Verskil is 3 Met 'n verskil van meer as 2,1 sal elektronoordrag plaasvind en die binding is ionies. (waarde van 2.1 hoef nie vir hierdie punt gegee te word nie) 4 marks/punte	LO2

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Table of Contents Physical Science 2014 Control Test

- 1. Understanding the eBook Physical Science 2014 Control Test
 - The Rise of Digital Reading Physical Science 2014 Control Test
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Science 2014 Control Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - $\circ \ \ Popular \ eBook \ Platforms$
 - \circ Features to Look for in an Physical Science 2014 Control Test
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Science 2014 Control Test
 - Personalized Recommendations
 - Physical Science 2014 Control Test User Reviews and Ratings
 - Physical Science 2014 Control Test and Bestseller Lists
- 5. Accessing Physical Science 2014 Control Test Free and Paid eBooks
 - Physical Science 2014 Control Test Public Domain eBooks

- Physical Science 2014 Control Test eBook Subscription Services
- Physical Science 2014 Control Test Budget-Friendly Options
- 6. Navigating Physical Science 2014 Control Test eBook Formats
 - o ePub, PDF, MOBI, and More
 - Physical Science 2014 Control Test Compatibility with Devices
 - Physical Science 2014 Control Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Physical Science 2014 Control Test
 - Highlighting and Note-Taking Physical Science 2014 Control Test
 - Interactive Elements Physical Science 2014 Control Test
- 8. Staying Engaged with Physical Science 2014 Control Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science 2014 Control Test
- 9. Balancing eBooks and Physical Books Physical Science 2014 Control Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science 2014 Control Test
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science 2014 Control Test
 - Setting Reading Goals Physical Science 2014 Control Test
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science 2014 Control Test
 - Fact-Checking eBook Content of Physical Science 2014 Control Test
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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